

Chapter One

Introduction

1.1 Brief History of Air Conditioning

The need for air-conditioning was first realized when it was found that certain products could be produced better in the right environment. With air-conditioning, there was no dependence of the product quality on the uncertainties of the weather and the factory sites were not limited to areas with a suitable climate. First to realize the importance of treatment of air was the textile industry. Industrial air conditioning later spread to various other areas such as manufacture of rayon and plastics, colour printing, pharmaceuticals, tobacco industry, manufacture, development, reproduction of photographic materials, electrical equipment, and many hygroscopic products.

Although sporadic attempts have been recorded for improving the ventilation and thermal conditions of indoor spaces, the first use of air-conditioning by mechanical means for human comfort was introduced in the United States of America at the turn of the twentieth century. There were few comfort cooling installations before 1920, but in the mid-twenties theatres all across the country were air conditioned to draw more customers. This was followed by the use of air conditioning in other commercial premises such as restaurants, night clubs, office blocks, etc.

In the 1930s, air conditioning was introduced in transportation systems such as trains, airplanes, buses and trolleys. After the Second World War, with the rising family income there was a boom in domestic air conditioning. This trend of use in the United States was subsequently repeated in the developing countries as more and more countries achieved higher standards of living.

1.2 System International Unites (SI Units)

SI or the International System of Units is the purest form and an extension and refinement of the traditional metric system. There are six basic SI units as given in Table 1.1. The units of other thermodynamic quantities may be derived from these basic units. The unit of temperature is kelvin which measures the absolute temperature given by

$$T = t + 273.15$$

where **t** is the Celsius temperature in °C.

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Table 1.1 Basic SI units

Quantity	Unit	Symbol
Length	meter	m
Mass	kilogram	kg
Time	second	s
Temperature	kelvin	K
Electrical current	ampere	A
Luminous intensity	candle	cd

1.2.1 Unit of Force

Force **F** is proportional to mass **m** and acceleration **a**, so that

$$F = C(m)(a)$$

Where **C** is a proportionality constant, and its value is taken as equal to the standard gravitational acceleration as (9.80665 m/s²). The SI unit of force is Newton denoted by the symbol **N**.

$$1 \text{ N} = 1 \text{ kg.m/s}^2$$

1.2.2 Units of Pressure

The SI unit of pressure **P** can also be derived from its definition as force per unit area

$$P = F / A \quad (\text{N/m}^2)$$

The unit is also called **Pascal** and is denoted by the symbol **Pa**.

Another common SI unit of pressure is **bar** which is equivalent to a pressure of 10⁵ N/m² or 0.1 MN/m² or 100 kN/m².

One standard atmosphere is given by

$$1 \text{ atm} = 1.01325 \text{ bar} = 760 \text{ mm Hg} = 760 \text{ torr}$$

$$1 \text{ torr} = 1 \text{ mm Hg} = 1/760 \text{ atm} = 133 \text{ N/m}^2$$

1.2.3 Unit of Energy (Work and Heat)

The SI unit of work is Newton meter (Nm) or Joule (J). Thus

$$1 \text{ Nm} = 1 \text{ J} = 1 \text{ kg m}^2/\text{s}^2$$

Since both heat and work are energy, the SI unit of heat is the same as the unit of work (J).

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1.2.4 Units of Power

The SI unit of power is Watt (W). It is defined as the rate of doing 1 Nm of work per second. Thus $1 \text{ W} = 1 \text{ J/s} = 1 \text{ Nm/s}$

It may also be noted that Watt also represents the electrical unit of work defined by

$$1 \text{ W} = 1(\text{Volt}) \times 1(\text{ampere}) = 1 \text{ J/s}$$

Further, the units of energy can be derived from those of power. Thus

$$1 \text{ J} = 1 \text{ W.s}$$

$$1 \text{ kWh} = 3600000 \text{ J} = 3600 \text{ kJ} = 860 \text{ kcal} = 3410 \text{ Btu}$$

1.2.5 Units of Enthalpy

The SI unit of specific enthalpy is J/kg or kJ/kg

$$1 \text{ kJ/kg} = 0.239 \text{ kcal/kg} = 0.42 \text{ Btu/lb}$$

1.2.6 Units of Entropy and Specific Heat

These are expressed as

$$1 \text{ kJ/kg.K} = 0.239 \text{ kcal/kg } ^\circ\text{C} = 0.239 \text{ Btu/lb } ^\circ\text{F}$$

1.2.7 Units of Refrigerating Capacity

The standard unit of refrigeration in vogue is ton refrigeration or simply (TR). It is equivalent to the production of cold at the rate at which heat is to be removed from one US tone of water at 32°F to freeze it to ice at 32°F in one day or 24 hours. Thus

$$1 \text{ TR} = (1 \times 2000 \text{ lb} \times 144 \text{ Btu/lb})/24 \text{ hr} = 12000 \text{ Btu/hr} = 200 \text{ Btu/min}$$

$$\text{Also, } 1 \text{ Btu} = 1.055 \text{ kJ}$$

$$1 \text{ TR} = 12000 \times 1.055 = 12660 \text{ kJ/hr} = 211 \text{ kJ/min} = 3.5167 \text{ kW}$$

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Moist Air Properties

In 1911 Willis H. Carrier made a significant contribution to the air-conditioning field when he published relations for moist air properties together with psychrometric chart. These formulas became fundamental to the industry.

In about 1945 Goff and Gratch published thermodynamic properties for moist air that were for many years the most accurate available. New formulations have recently been developed at the National Bureau of Standards. The properties based on these formulations are the basis for the thermodynamic properties of moist air given in the ASHRAE Handbook, Fundamentals.

Moist Air and The Standard Atmosphere

Atmospheric air is a mixture of many gases plus water vapor and countless pollutants. Aside from the pollutants, which may vary considerably from place to place, the composition of the dry air alone, is relatively constant, varying slightly with time, location, and altitude. The ASHRAE Handbook, Fundamentals gives the following approximate composition of dry air by volume

Nitrogen	0.78084
Oxygen	0.20948
Argon	0.00934
Carbon dioxide	0.00031
Neon, helium, sulfur dioxide, hydrogen, and other minor gases	0.00003

Based on the composition of air in Table 2.1, the molecular weight (M_a) of dry air is:

$$\begin{aligned} M_a &= M_{O_2} \times F_{O_2} + M_N \times F_N + M_A \times F_A + M_{CO_2} \times F_{CO_2} \dots\dots\dots(2.1) \\ &= 32 \times 0.209 + 28.016 \times 0.7809 + 39.94 \times 0.0093 + 44.014 \times 0.0003 \\ &= 28.965 \end{aligned}$$

Table 2.1

Constituent	Molecular Weight	Volume Fraction
Oxygen	32.000	0.2095
Nitrogen	28.016	0.7809
Argon	39.944	0.0093
Carbon dioxide	44.014	0.0003

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And the gas constant of the air (R_a) can be calculate from equation (2.2)

$$R_a = \frac{R}{M_a} = \frac{8314}{28.965} = 287 \text{ J/kg.K} \dots\dots\dots (2.2)$$

Where:

R: Universal gas constant

R_a : gas (air) constant

The molecular weight of water is 18.015 and the gas constant for water vapor is:

$$R_v = \frac{8314}{18.015} = 461 \text{ J/kg.K}$$

The universal gas law (perfect gas) is:

$$P V = m R T \dots\dots\dots (2.3)$$

$$P = \rho R T$$

Standard sea level density compute using Eq. (2.3) with the standard temperature and pressure is 1.225 kg/m^3 of dry air.

Atmospheric pressure may be computed as a function of elevation by the following relation:

$$P = a + b H \dots\dots\dots(2.4)$$

Where the constant **a** and **b** are:

$$a = 101.325 \quad , \quad b = - 0.01153 \quad \text{when } H \leq 1220 \text{ m}$$

$$a = 29.42 \quad , \quad b = - 0.0009 \quad \text{when } H > 1200 \text{ m}$$

H : elevation above sea level in meter.

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TABLE A-2 Properties of Saturated Water (Liquid–Vapor): Temperature Table

Temp. °C	Press. bar	Specific Volume m ³ /kg		Internal Energy kJ/kg		Enthalpy kJ/kg			Entropy kJ/kg · K		Temp. °C
		Sat. Liquid $v_f \times 10^3$	Sat. Vapor v_g	Sat. Liquid u_f	Sat. Vapor u_g	Sat. Liquid h_f	Evap. h_{fg}	Sat. Vapor h_g	Sat. Liquid s_f	Sat. Vapor s_g	
.01	0.00611	1.0002	206.136	0.00	2375.3	0.01	2501.3	2501.4	0.0000	9.1562	.01
4	0.00813	1.0001	157.232	16.77	2380.9	16.78	2491.9	2508.7	0.0610	9.0514	4
5	0.00872	1.0001	147.120	20.97	2382.3	20.98	2489.6	2510.6	0.0761	9.0257	5
6	0.00935	1.0001	137.734	25.19	2383.6	25.20	2487.2	2512.4	0.0912	9.0003	6
8	0.01072	1.0002	120.917	33.59	2386.4	33.60	2482.5	2516.1	0.1212	8.9501	8
10	0.01228	1.0004	106.379	42.00	2389.2	42.01	2477.7	2519.8	0.1510	8.9008	10
11	0.01312	1.0004	99.857	46.20	2390.5	46.20	2475.4	2521.6	0.1658	8.8765	11
12	0.01402	1.0005	93.784	50.41	2391.9	50.41	2473.0	2523.4	0.1806	8.8524	12
13	0.01497	1.0007	88.124	54.60	2393.3	54.60	2470.7	2525.3	0.1953	8.8285	13
14	0.01598	1.0008	82.848	58.79	2394.7	58.80	2468.3	2527.1	0.2099	8.8048	14
15	0.01705	1.0009	77.926	62.99	2396.1	62.99	2465.9	2528.9	0.2245	8.7814	15
16	0.01818	1.0011	73.333	67.18	2397.4	67.19	2463.6	2530.8	0.2390	8.7582	16
17	0.01938	1.0012	69.044	71.38	2398.8	71.38	2461.2	2532.6	0.2535	8.7351	17
18	0.02064	1.0014	65.038	75.57	2400.2	75.58	2458.8	2534.4	0.2679	8.7123	18
19	0.02198	1.0016	61.293	79.76	2401.6	79.77	2456.5	2536.2	0.2823	8.6897	19
20	0.02339	1.0018	57.791	83.95	2402.9	83.96	2454.1	2538.1	0.2966	8.6672	20
21	0.02487	1.0020	54.514	88.14	2404.3	88.14	2451.8	2539.9	0.3109	8.6450	21
22	0.02645	1.0022	51.447	92.32	2405.7	92.33	2449.4	2541.7	0.3251	8.6229	22
23	0.02810	1.0024	48.574	96.51	2407.0	96.52	2447.0	2543.5	0.3393	8.6011	23
24	0.02985	1.0027	45.883	100.70	2408.4	100.70	2444.7	2545.4	0.3534	8.5794	24
25	0.03169	1.0029	43.360	104.88	2409.8	104.89	2442.3	2547.2	0.3674	8.5580	25
26	0.03363	1.0032	40.994	109.06	2411.1	109.07	2439.9	2549.0	0.3814	8.5367	26
27	0.03567	1.0035	38.774	113.25	2412.5	113.25	2437.6	2550.8	0.3954	8.5156	27
28	0.03782	1.0037	36.690	117.42	2413.9	117.43	2435.2	2552.6	0.4093	8.4946	28
29	0.04008	1.0040	34.733	121.60	2415.2	121.61	2432.8	2554.5	0.4231	8.4739	29
30	0.04246	1.0043	32.894	125.78	2416.6	125.79	2430.5	2556.3	0.4369	8.4533	30
31	0.04496	1.0046	31.165	129.96	2418.0	129.97	2428.1	2558.1	0.4507	8.4329	31
32	0.04759	1.0050	29.540	134.14	2419.3	134.15	2425.7	2559.9	0.4644	8.4127	32
33	0.05034	1.0053	28.011	138.32	2420.7	138.33	2423.4	2561.7	0.4781	8.3927	33
34	0.05324	1.0056	26.571	142.50	2422.0	142.50	2421.0	2563.5	0.4917	8.3728	34
35	0.05628	1.0060	25.216	146.67	2423.4	146.68	2418.6	2565.3	0.5053	8.3531	35
36	0.05947	1.0063	23.940	150.85	2424.7	150.86	2416.2	2567.1	0.5188	8.3336	36
38	0.06632	1.0071	21.602	159.20	2427.4	159.21	2411.5	2570.7	0.5458	8.2950	38
40	0.07384	1.0078	19.523	167.56	2430.1	167.57	2406.7	2574.3	0.5725	8.2570	40
45	0.09593	1.0099	15.258	188.44	2436.8	188.45	2394.8	2583.2	0.6387	8.1648	45

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TABLE A-2 (Continued)

Temp. °C	Press. bar	Specific Volume m ³ /kg		Internal Energy kJ/kg		Enthalpy kJ/kg			Entropy kJ/kg · K		Temp. °C
		Sat. Liquid $v_f \times 10^3$	Sat. Vapor v_g	Sat. Liquid u_f	Sat. Vapor u_g	Sat. Liquid h_f	Evap. h_{fg}	Sat. Vapor h_g	Sat. Liquid s_f	Sat. Vapor s_g	
50	.1235	1.0121	12.032	209.32	2443.5	209.33	2382.7	2592.1	.7038	8.0763	50
55	.1576	1.0146	9.568	230.21	2450.1	230.23	2370.7	2600.9	.7679	7.9913	55
60	.1994	1.0172	7.671	251.11	2456.6	251.13	2358.5	2609.6	.8312	7.9096	60
65	.2503	1.0199	6.197	272.02	2463.1	272.06	2346.2	2618.3	.8935	7.8310	65
70	.3119	1.0228	5.042	292.95	2469.6	292.98	2333.8	2626.8	.9549	7.7553	70
75	.3858	1.0259	4.131	313.90	2475.9	313.93	2321.4	2635.3	1.0155	7.6824	75
80	.4739	1.0291	3.407	334.86	2482.2	334.91	2308.8	2643.7	1.0753	7.6122	80
85	.5783	1.0325	2.828	355.84	2488.4	355.90	2296.0	2651.9	1.1343	7.5445	85
90	.7014	1.0360	2.361	376.85	2494.5	376.92	2283.2	2660.1	1.1925	7.4791	90
95	.8455	1.0397	1.982	397.88	2500.6	397.96	2270.2	2668.1	1.2500	7.4159	95
100	1.014	1.0435	1.673	418.94	2506.5	419.04	2257.0	2676.1	1.3069	7.3549	100
110	1.433	1.0516	1.210	461.14	2518.1	461.30	2230.2	2691.5	1.4185	7.2387	110
120	1.985	1.0603	0.8919	503.50	2529.3	503.71	2202.6	2706.3	1.5276	7.1296	120
130	2.701	1.0697	0.6685	546.02	2539.9	546.31	2174.2	2720.5	1.6344	7.0269	130
140	3.613	1.0797	0.5089	588.74	2550.0	589.13	2144.7	2733.9	1.7391	6.9299	140
150	4.758	1.0905	0.3928	631.68	2559.5	632.20	2114.3	2746.5	1.8418	6.8379	150
160	6.178	1.1020	0.3071	674.86	2568.4	675.55	2082.6	2758.1	1.9427	6.7502	160
170	7.917	1.1143	0.2428	718.33	2576.5	719.21	2049.5	2768.7	2.0419	6.6663	170
180	10.02	1.1274	0.1941	762.09	2583.7	763.22	2015.0	2778.2	2.1396	6.5857	180
190	12.54	1.1414	0.1565	806.19	2590.0	807.62	1978.8	2786.4	2.2359	6.5079	190
200	15.54	1.1565	0.1274	850.65	2595.3	852.45	1940.7	2793.2	2.3309	6.4323	200
210	19.06	1.1726	0.1044	895.53	2599.5	897.76	1900.7	2798.5	2.4248	6.3585	210
220	23.18	1.1900	0.08619	940.87	2602.4	943.62	1858.5	2802.1	2.5178	6.2861	220
230	27.95	1.2088	0.07158	986.74	2603.9	990.12	1813.8	2804.0	2.6099	6.2146	230
240	33.44	1.2291	0.05976	1033.2	2604.0	1037.3	1766.5	2803.8	2.7015	6.1437	240
250	39.73	1.2512	0.05013	1080.4	2602.4	1085.4	1716.2	2801.5	2.7927	6.0730	250
260	46.88	1.2755	0.04221	1128.4	2599.0	1134.4	1662.5	2796.6	2.8838	6.0019	260
270	54.99	1.3023	0.03564	1177.4	2593.7	1184.5	1605.2	2789.7	2.9751	5.9301	270
280	64.12	1.3321	0.03017	1227.5	2586.1	1236.0	1543.6	2779.6	3.0668	5.8571	280
290	74.36	1.3656	0.02557	1278.9	2576.0	1289.1	1477.1	2766.2	3.1594	5.7821	290
300	85.81	1.4036	0.02167	1332.0	2563.0	1344.0	1404.9	2749.0	3.2534	5.7045	300
320	112.7	1.4988	0.01549	1444.6	2525.5	1461.5	1238.6	2700.1	3.4480	5.5362	320
340	145.9	1.6379	0.01080	1570.3	2464.6	1594.2	1027.9	2622.0	3.6594	5.3357	340
360	186.5	1.8925	0.006945	1725.2	2351.5	1760.5	720.5	2481.0	3.9147	5.0526	360
374.14	220.9	3.155	0.003155	2029.6	2029.6	2099.3	0	2099.3	4.4298	4.4298	374.14

H₂O

Source: Tables A-2 through A-5 are extracted from J. H. Keenan, F. G. Keyes, P. G. Hill, and J. G. Moore, *Steam Tables*, Wiley, New York, 1969.

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TABLE A-3 Properties of Saturated Water (Liquid–Vapor): Pressure Table

H ₂ O	Press. bar	Temp. °C	Specific Volume m ³ /kg		Internal Energy kJ/kg		Enthalpy kJ/kg			Entropy kJ/kg · K		Press. bar
			Sat. Liquid $v_f \times 10^3$	Sat. Vapor v_g	Sat. Liquid u_f	Sat. Vapor u_g	Sat. Liquid h_f	Evap. h_{fg}	Sat. Vapor h_g	Sat. Liquid s_f	Sat. Vapor s_g	
0.04	28.96	1.0040	34.800	121.45	2415.2	121.46	2432.9	2554.4	0.4226	8.4746	0.04	
0.06	36.16	1.0064	23.739	151.53	2425.0	151.53	2415.9	2567.4	0.5210	8.3304	0.06	
0.08	41.51	1.0084	18.103	173.87	2432.2	173.88	2403.1	2577.0	0.5926	8.2287	0.08	
0.10	45.81	1.0102	14.674	191.82	2437.9	191.83	2392.8	2584.7	0.6493	8.1502	0.10	
0.20	60.06	1.0172	7.649	251.38	2456.7	251.40	2358.3	2609.7	0.8320	7.9085	0.20	
0.30	69.10	1.0223	5.229	289.20	2468.4	289.23	2336.1	2625.3	0.9439	7.7686	0.30	
0.40	75.87	1.0265	3.993	317.53	2477.0	317.58	2319.2	2636.8	1.0259	7.6700	0.40	
0.50	81.33	1.0300	3.240	340.44	2483.9	340.49	2305.4	2645.9	1.0910	7.5939	0.50	
0.60	85.94	1.0331	2.732	359.79	2489.6	359.86	2293.6	2653.5	1.1453	7.5320	0.60	
0.70	89.95	1.0360	2.365	376.63	2494.5	376.70	2283.3	2660.0	1.1919	7.4797	0.70	
0.80	93.50	1.0380	2.087	391.58	2498.8	391.66	2274.1	2665.8	1.2329	7.4346	0.80	
0.90	96.71	1.0410	1.869	405.06	2502.6	405.15	2265.7	2670.9	1.2695	7.3949	0.90	
1.00	99.63	1.0432	1.694	417.36	2506.1	417.46	2258.0	2675.5	1.3026	7.3594	1.00	
1.50	111.4	1.0528	1.159	466.94	2519.7	467.11	2226.5	2693.6	1.4336	7.2233	1.50	
2.00	120.2	1.0605	0.8857	504.49	2529.5	504.70	2201.9	2706.7	1.5301	7.1271	2.00	
2.50	127.4	1.0672	0.7187	535.10	2537.2	535.37	2181.5	2716.9	1.6072	7.0527	2.50	
3.00	133.6	1.0732	0.6058	561.15	2543.6	561.47	2163.8	2725.3	1.6718	6.9919	3.00	
3.50	138.9	1.0786	0.5243	583.95	2546.9	584.33	2148.1	2732.4	1.7275	6.9405	3.50	
4.00	143.6	1.0836	0.4625	604.31	2553.6	604.74	2133.8	2738.6	1.7766	6.8959	4.00	
4.50	147.9	1.0882	0.4140	622.25	2557.6	623.25	2120.7	2743.9	1.8207	6.8565	4.50	
5.00	151.9	1.0926	0.3749	639.68	2561.2	640.23	2108.5	2748.7	1.8607	6.8212	5.00	
6.00	158.9	1.1006	0.3157	669.90	2567.4	670.56	2086.3	2756.8	1.9312	6.7600	6.00	
7.00	165.0	1.1080	0.2729	696.44	2572.5	697.22	2066.3	2763.5	1.9922	6.7080	7.00	
8.00	170.4	1.1148	0.2404	720.22	2576.8	721.11	2048.0	2769.1	2.0462	6.6628	8.00	
9.00	175.4	1.1212	0.2150	741.83	2580.5	742.83	2031.1	2773.9	2.0946	6.6226	9.00	
10.0	179.9	1.1273	0.1944	761.68	2583.6	762.81	2015.3	2778.1	2.1387	6.5863	10.0	
15.0	198.3	1.1539	0.1318	843.16	2594.5	844.84	1947.3	2792.2	2.3150	6.4448	15.0	
20.0	212.4	1.1767	0.09963	906.44	2600.3	908.79	1890.7	2799.5	2.4474	6.3409	20.0	
25.0	224.0	1.1973	0.07998	959.11	2603.1	962.11	1841.0	2803.1	2.5547	6.2575	25.0	
30.0	233.9	1.2165	0.06668	1004.8	2604.1	1008.4	1795.7	2804.2	2.6457	6.1869	30.0	
35.0	242.6	1.2347	0.05707	1045.4	2603.7	1049.8	1753.7	2803.4	2.7253	6.1253	35.0	
40.0	250.4	1.2522	0.04978	1082.3	2602.3	1087.3	1714.1	2801.4	2.7964	6.0701	40.0	
45.0	257.5	1.2692	0.04406	1116.2	2600.1	1121.9	1676.4	2798.3	2.8610	6.0199	45.0	
50.0	264.0	1.2859	0.03944	1147.8	2597.1	1154.2	1640.1	2794.3	2.9202	5.9734	50.0	
60.0	275.6	1.3187	0.03244	1205.4	2589.7	1213.4	1571.0	2784.3	3.0267	5.8892	60.0	
70.0	285.9	1.3513	0.02737	1257.6	2580.5	1267.0	1505.1	2772.1	3.1211	5.8133	70.0	
80.0	295.1	1.3842	0.02352	1305.6	2569.8	1316.6	1441.3	2758.0	3.2068	5.7432	80.0	
90.0	303.4	1.4178	0.02048	1350.5	2557.8	1363.3	1378.9	2742.1	3.2858	5.6772	90.0	
100.	311.1	1.4524	0.01803	1393.0	2544.4	1407.6	1317.1	2724.7	3.3596	5.6141	100.	
110.	318.2	1.4886	0.01599	1433.7	2529.8	1450.1	1255.5	2705.6	3.4295	5.5527	110.	

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Tables in SI Units 723

TABLE A-3 (Continued)

Press. bar	Temp. °C	Specific Volume m ³ /kg		Internal Energy kJ/kg		Enthalpy kJ/kg			Entropy kJ/kg · K		Press. bar
		Sat. Liquid $v_f \times 10^3$	Sat. Vapor v_g	Sat. Liquid u_f	Sat. Vapor u_g	Sat. Liquid h_f	Evap. h_{fg}	Sat. Vapor h_g	Sat. Liquid s_f	Sat. Vapor s_g	
120.	324.8	1.5267	0.01426	1473.0	2513.7	1491.3	1193.6	2684.9	3.4962	5.4924	120.
130.	330.9	1.5671	0.01278	1511.1	2496.1	1531.5	1130.7	2662.2	3.5606	5.4323	130.
140.	336.8	1.6107	0.01149	1548.6	2476.8	1571.1	1066.5	2637.6	3.6232	5.3717	140.
150.	342.2	1.6581	0.01034	1585.6	2455.5	1610.5	1000.0	2610.5	3.6848	5.3098	150.
160.	347.4	1.7107	0.009306	1622.7	2431.7	1650.1	930.6	2580.6	3.7461	5.2455	160.
170.	352.4	1.7702	0.008364	1660.2	2405.0	1690.3	856.9	2547.2	3.8079	5.1777	170.
180.	357.1	1.8397	0.007489	1698.9	2374.3	1732.0	777.1	2509.1	3.8715	5.1044	180.
190.	361.5	1.9243	0.006657	1739.9	2338.1	1776.5	688.0	2464.5	3.9388	5.0228	190.
200.	365.8	2.036	0.005834	1785.6	2293.0	1826.3	583.4	2409.7	4.0139	4.9269	200.
220.9	374.1	3.155	0.003155	2029.6	2029.6	2099.3	0	2099.3	4.4298	4.4298	220.9

H₂O

Chapter Two

Fundamental Parameters

The Gibbs Dalton law for a mixture of perfect gases states that the mixture pressure is equal to the sum of the partial pressure of the constituents

$$P = p_1 + p_2 + p_3 + \dots \quad (2.5)$$

For moist air

$$P = p_{N_2} + p_{O_2} + p_{CO_2} + p_A + p_v \quad (2.6)$$

Because the various constituents of the dry air may be considered to be one gas,

$$P_B = p_a + p_v \quad (2.7)$$

Where:

P_B : Barometric pressure (Atmospheric pressure)

P_a : dry air pressure

P_v : water vapor pressure

Example:

Saturated air at 26°C and atmospheric pressure 101325 N/m². Find the partial pressure for each of the dry air and the water vapor?

Solution:

From table A-2 at temperature 26°C, the saturated pressure of the water vapor is

$P_v = 3363 \text{ Pa (N/m}^2\text{)}$, and from Eqn. (2.7)

$$P_B = p_a + p_v \quad \text{or} \quad p_a = P_B - p_v \\ = 101325 - 3363 = 97962 \text{ Pa} = 97.962 \text{ kPa}$$

Vapor partial pressure in unsaturated air

These cases are the most common in nature, the vapor pressure can be calculated from empirical formula:

$$P_v = P_{wss} - P_B A (t_d - t_w) \quad (2.8)$$

Where:

P_v : partial pressure of water vapor

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P_{wss} : saturated pressure of vapor at wet bulb temperature (t_w)

P_B : barometric pressure

t_d : dry bulb temperature

t_w : wet bulb temperature

A: constant

$$A = 6.66 \times 10^{-4} \quad \text{if } t_w \geq 0^\circ\text{C}$$

$$A = 5.94 \times 10^{-4} \quad \text{if } t_w < 0^\circ\text{C}$$

Example:

Calculate the water vapor pressure in wet air under 20°C dry bulb and 15°C wet bulb temperature and $P_B = 950$ mbar?

Solution:

$$P_v = P_{wss} - P_B A (t_d - t_w)$$

From steam table A-2 at $t_w = 15^\circ\text{C}$

$$P_{wss} = 1.705 \text{ kPa}$$

$$P_v = 1.705 - 95 \times 6.66 \times 10^{-4} (20 - 15) = 1.388 \text{ kPa}$$

Moisture content (W)

Sometimes called the specific humidity or humidity ratio, it is the ratio of the mass of water vapor (m_v) to the mass of the dry air (m_a) in the mixture:

$$W = m_v / m_a \dots\dots\dots (2.8)$$

To find the values of (m_v) and (m_a) using equation (2.3)

$$m_v = P_v V_v / R_v T_v, \quad m_a = P_a V_a / R_a T_a$$

but $W = m_v / m_a$

and $V_v = V_a, \quad T_v = T_a$

$$W = (P_v/R_v) / (P_a/R_a) = (P_v/P_a) \cdot (R_a/R_v) = 0.622 P_v / P_a = 0.622 P_v / (P_B - P_v) \dots\dots\dots (2.9)$$

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Example:

Air at 20°C dry bulb and 15°C wet bulb and the barometric pressure is 95 kPa. Calculate the moisture content of the air and the density of the vapor.

Solution:

From steam table A-2 at $t_w = 15^\circ\text{C}$

$$P_{wss} = 1.705 \text{ kPa}$$

$$P_v = P_{wss} - P_B \cdot A \cdot (t_d - t_w)$$

$$= 1.705 - 95 \times 6.66 \times 10^{-4} (20 - 15) = 1.389 \text{ kPa}$$

$$W = 0.622 (P_v) / (P_B - P_v) = 0.622 (1.389) / (95 - 1.389) = 0.00923 \text{ kg/kg dry air}$$

Density of the vapor, $P_v = \rho_v \cdot R_v \cdot T_v \rightarrow \rho_v = (1389) / (461)(20+273) = 0.01028 \text{ kg/m}^3$

Relative humidity (Φ)

Is the ratio of the mole fraction of the water vapor (X_v) in a mixture to the mole fraction of the water vapor in a saturated mixture (X_s) at the same temperature and pressure.

$$\Phi = [X_v / X_s]_{t, p} \dots\dots\dots (2.10)$$

For a mixture of perfect gases, the mole fraction is equal to the partial pressure ratio of each constituent. The mole fraction of water vapor is:

$$X_v = P_v / P$$

Then; $\Phi = (P_v / P) / (P_{ss} / P) = P_v / P_{ss} \dots\dots\dots (2.11)$

Since the temperature of the dry air and the water vapor are assumed to be the same in the mixture;

$$\Phi = (P_v / R_v \cdot T) / (P_{ss} / R_v \cdot T) = [\rho_v / \rho_{ss}]_{t, p} \dots\dots\dots (2.12)$$

By using the perfect gas law we can derive a relation between the relative humidity (Φ) and the moisture content (W). Combining Eqns. (2.9) and (2.11)

$$\Phi = (W \cdot P_B) / (0.622 + W) P_{ss} \dots\dots\dots (2.13)$$

Example:

Air at 24°C and 40% RH, and $P_B = 92 \text{ kPa}$. Find the vapor density (ρ_v) and vapor pressure (P_v).

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Solution:

From Table A-2 at $t_d=24^\circ\text{C}$, $v = 45.883 \text{ m}^3/\text{kg}$

$$\rho_{ss} = 1/v = 1/45.883 = 0.02179 \text{ kg/m}^3$$

$$\Phi = [\rho_v / \rho_{ss}]_{t,p} \rightarrow \rho_v = \Phi \cdot \rho_{ss} = 0.4 \times 0.02179 = 0.008716 \text{ kg/m}^3$$

$$P_v = \rho_v R_v T = 0.008716 \times 461 \times (24+273) = 1193.4 \text{ Pa}$$

Or by other way;

$$\Phi = [\rho_v / \rho_{ss}]_{t,p}, \text{ from Table A-2 at } t=24^\circ\text{C} \rightarrow P_{ss} = 2985 \text{ Pa}$$

$$P_v = 0.4 \times 2985 = 1194 \text{ Pa}$$

Degree of saturation (μ)

Is the ratio of the moisture content (W) to the moisture content of a saturated mixture (W_{ss}) at the same temperature and pressure.

$$\mu = [W/W_{ss}]_{t,p} \dots\dots\dots (2.14)$$

$$W = m_s/m_a = (P_s/(P_B - P_s)) \cdot (R_a/R_s)$$

$$W_{ss} = (P_{ss}/(P_B - P_{ss})) \cdot (R_a/R_s)$$

$$\mu = (P_s/(P_B - P_s)) \cdot ((P_B - P_{ss})/P_{ss}) \cdot (R_a/R_s)(R_s/R_a) = (P_s/P_{ss}) \cdot (P_B - P_{ss})/(P_B - P_s)$$

$$= (P_s/P_{ss}) \cdot (1 - P_{ss}/P_B)/(1 - P_s/P_B) = \Phi [(1 - P_{ss}/P_B)/(1 - \Phi \cdot P_{ss}/P_B)] \dots\dots\dots (2.15)$$

Example:

Moist air at 40°C DBT, 30°C WBT and 101 kPa barometric pressure, calculate for the air:

- a) Relative humidity (Φ)
- b) Moisture content (W)
- c) Degree of saturation (μ)

Solution:

$$P_v = P_{wss} - P_B \cdot A \cdot (t_d - t_w), \text{ from Table A-2}$$

$$P_{wss} = 4.246 \text{ kPa at } 30^\circ\text{C}, P_{dss} = 7.384 \text{ kPa at } 40^\circ\text{C}$$

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$$P_v = 4.246 - 101 \times 6.66 \times 10^{-4} (40 - 30) = 3.57 \text{ kPa}$$

$$\text{a) } \Phi = P_v/P_{dss} \times 100\% = (3.57)/(7.384) \times 100\% = 48.4\%$$

$$\text{b) } W = 0.622 P_v / (P_B - P_v) = 0.622(3.57)/(101 - 3.57) = 0.0228 \text{ kg/kg}_{d.a}$$

$$\text{c) } \mu = [W/W_{ss}]_{t,p}$$

$$W_{ss} = 0.622(7.384)/(101 - 7.384) = 0.049 \text{ kg/kg}_{d.a}$$

$$\mu = 0.0228/0.049 = 0.465 = 46.5\%$$

Or:

$$\begin{aligned} \mu &= \Phi [(1 - P_{ss}/P_B) / (1 - \Phi \cdot P_{ss}/P_B)] = (0.484) [1 - (7.384/101)] / [1 - 0.484(7.384/101)] \\ &= 46.5\% \end{aligned}$$

Dew point temperature (t_{dp})

This is defined as the temperature of saturated air which has the same vapor pressure as the moist air under consideration. It may also be stated as a mixture is cooled at constant pressure, the temperature at which condensation first begins is the dew point.

Example:

Find the dew point temperature for air at 20°C DBT and 15°C WBT and atmospheric pressure is 95 kPa.

Solution:

From Table A-2 at $t_w = 15^\circ\text{C}$, $P_{wss} = 1.705 \text{ kPa}$

$$P_v = P_{wss} - P_B \cdot A \cdot (t_d - t_w)$$

$$= 1.705 - 95 \times 6.66 \times 10^{-4} (20 - 15) = 1.388 \text{ kPa}$$

From Table A-2, we find the saturation temperature that corresponds to P_v

$$t_{sat} = t_{dp} = 11.84^\circ\text{C}$$

Specific volume (v)

This is defined as the volume of one kg of moist air, it can be calculated by three ways:

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1. Used of dry air mass and partial pressure
2. Used of vapor mass and partial pressure
3. Used of mass of mixture and its total pressure.

Example:

Calculate the specific volume (v) of air at 35°C DBT and 25°C WBT at a barometric pressure of 95 kPa.

Solution:

$$P_a V_a = m_a R_a T_a \rightarrow v_a = R_a T_a / P_a$$

$$\text{But; } P_a = P_B - P_v$$

$$P_{wss} = 3.169 \text{ kPa at } t = 25^\circ\text{C}$$

$$P_v = 3.169 - 95 \times 6.66 \times 10^{-4} (35 - 25) = 2.5363 \text{ kPa}$$

$$\text{Then } P_a = 95 - 2.5363 = 92.46 \text{ kPa}$$

$$v_a = (0.287)(35 + 273)/(92.46) = 0.956 \text{ m}^3/\text{kg}$$

Or by other way;

$$W = 0.622 (P_v)/(P_B - P_v) = 0.622(2.5363)/(95 - 2.5363) = 0.01706 \text{ kg/kg}_{d.a}$$

So for 1 kg of dry air there is 0.01706 kg of steam ; $m_s = 0.01706 \text{ kg}$

$$v_v = (0.01706)(0.461)(35 + 273)/(2.5363) = 0.956 \text{ m}^3/\text{kg}$$

Enthalpy (h)

If a heat exchange occurs at constant pressure, as well as a change in internal energy taking place, work may be done.

This leads to a definition of enthalpy, H:

$$H = U + pV$$

The equation is strictly true for a pure gas of mass m , pressure p , and volume V . However it may be applied without appreciable error to the mixtures of gases associated with air conditioning.

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The enthalpy, h , used in psychrometry is the specific enthalpy of moist air, expressed in kJ/kg dry air, defined by the equation:

$$h = h_a + W h_v \dots\dots\dots (2.16)$$

where ; h_a is the enthalpy of dry air, h_v is the enthalpy of water vapor, both expressed in kJ kg^{-1} , and W is the moisture content.

An approximate equation for the enthalpy of dry air over the range 0°C to 60°C is, however $h_a = 1.007 t_d - 0.026 \dots\dots\dots (2.17)$

and for lower temperatures, down to -10°C , the approximate equation is

$$h_a = 1.005 t_d \dots\dots\dots (2.18)$$

For purposes of approximate calculation, we may assume that, in the range 0°C to 60°C , the vapor is generated from liquid water at 0°C and that the specific heat of superheated steam is a constant. The following equation can then be used for the enthalpy of water vapor:

$$h_v = 2501 + 1.84 t_d \dots\dots\dots (2.19)$$

Equations (2.17) and (2.19) can now be combined, as typified by equation (2.16), to give an approximate expression for the enthalpy of humid air at a barometric pressure of 101.325 kPa:

$$h = (1.007 t_d - 0.026) + W(2501 + 1.84 t_d) \dots\dots\dots (2.20)$$

Example:

Compute the enthalpy of saturated air at 15°C and standard atmospheric pressure.

Solution:

Because the air is saturated, $W = W_s$

$$W_s = 0.622(P_{vs})/(P_B - P_{vs}) \quad \text{from Table A-2 at } t=15^\circ\text{C, } P_{vs} = 1705 \text{ Pa}$$

$$W_s = 0.622(1705)/(101325 - 1705) = 0.01065 \text{ kg/kg}_{d.a}$$

$$h = (1.007 t_d - 0.026) + W(2501 + 1.84 t_d)$$

$$= (1.007 \times 15 - 0.026) + 0.01065(2501 + 1.84 \times 15) = 42 \text{ kJ/kg}$$

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Example:

Compute the enthalpy of moist air at 20°C DBT and 15°C WBT and standard atmospheric pressure.

Solution:

$$P_v = P_{wss} - P_B \cdot A \cdot (t_d - t_w) \\ = 1.705 - 101.325 \times 6.66 \times 10^{-4} (20 - 15) = 1.3676 \text{ kPa}$$

$$W_s = 0.622(P_{vs})/(P_B - P_{vs}) \\ = 0.622(1.3676)/(101.325 - 1.3676) = 0.00843 \text{ kg/kg}_{d.a}$$

$$h = (1.007 \times 20 - 0.026) + 0.00843(2501 + 1.84 \times 20) = 41.508 \text{ kJ/kg}$$

Example:

Moist air at 42°C DBT, 26°C WBT and 100 kPa barometric pressure. Calculate:

- a) vapor pressure b) relative humidity c) moisture content d) specific enthalpy
e) specific volume f) dew point temperature g) degree of saturation

Solution:

From Table A-2

$$P_{wss} = 3.363 \text{ kPa} \quad \text{at } t_w = 26^\circ\text{C}$$

$$P_{dss} = 8.268 \text{ kPa} \quad \text{at } t_d = 42^\circ\text{C}$$

a) $P_v = 3.363 - 100 \times 6.66 \times 10^{-4} (42 - 26) = 2.2974 \text{ kPa}$

b) $\Phi = P_v/P_{dss} = 2.2974/8.268 = 0.2779 = 27.79\%$

c) $W = 0.622 (2.2974)/(100 - 2.2974) = 0.0146 \text{ kg/kg}_{d.a}$

d) $h = (1.007 t_d - 0.026) + W(2501 + 1.84 t_d) \\ = (1.007 \times 42 - 0.026) + 0.0146(2501 + 1.84 \times 42) = 79.91 \text{ kJ/kg}$

e) $v_a = R_a T_a / P_a = (0.287)(42 + 273)/(100 - 2.2974) = 0.9253 \text{ m}^3/\text{kg}$

f) the dew point temperature is the temperature corresponds to P_v

$$t_{dp} = 19.7^\circ\text{C}$$

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$$g) \quad \mu = \Phi \left[\frac{(1 - P_{ss}/P_B)}{(1 - \Phi \cdot P_{ss}/P_B)} \right] = 0.2779 \left[\frac{(1 - 8.268/100)}{(1 - 0.2779 \times 8.268/100)} \right] \\ = 0.26092$$

Example:

In a check on an air-conditioning system, a room was maintained at 22°C DBT and 53% RH, by air supply at 1.6 m³/s with 10°C DBT and 9°C WBT. The barometric pressure was 920 mbar. Find:

a) sensible heat gain b) latent heat gain c) total heat gain

Solution:

From Table A-2 at $t_w = 9^\circ\text{C}$, $P_{wss} = 1.15 \text{ kPa}$

$$P_v = 1.15 - 92 \times 6.66 \times 10^{-4} (10 - 9) = 1.0887 \text{ kPa}$$

$$P_a = 92 - 1.0887 = 90.91 \text{ kPa}$$

$$m'_a = (P_a V_a) / (R_a T_a) = (90.91 \times 1.6) / (0.287 \times 295) = 1.718 \text{ kg/s}$$

$$h_{a1} = 1.007 \times 22 - 0.026 = 22.128 \text{ kJ/kg}$$

$$h_{a2} = 1.007 \times 10 - 0.026 = 10.044 \text{ kJ/kg}$$

$$Q_s = m'_a (h_{a1} - h_{a2}) = 1.718 (22.128 - 10.044) = 20.7 \text{ kW}$$

Also, from Table A-2 at $t = 22^\circ\text{C}$ $P_{dss} = 2.645 \text{ kPa}$

$$\Phi = P_v / P_{dss} \rightarrow P_v = 0.53 \times 2.645 = 1.4 \text{ kPa}$$

$$W_1 = 0.622(1.4) / (92 - 1.4) = 0.00961 \text{ kg/kg}_{d.a}$$

$$W_2 = 0.622(1.0887) / (92 - 1.0887) = 0.00745 \text{ kg/kg}_{d.a}$$

$$h_{v1} = 2501 + 1.84 \times 22 = 2541.48 \text{ kJ/kg}$$

$$h_{v2} = 2501 + 1.84 \times 10 = 2519.4 \text{ kJ/kg}$$

$$Q_L = m'_a (W_1 \cdot h_{v1} - W_2 \cdot h_{v2}) = 1.718 (0.00961 \times 2541.48 - 0.00745 \times 2519.4) = 9.714 \text{ kW}$$

$$Q_{\text{Tot}} = Q_s + Q_L = 20.7 + 9.714 = 30.414 \text{ kW}$$

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Adiabatic saturation and thermodynamic wet bulb temperature

Adiabatic saturation temperature is defined as that temperature at which water, by evaporating into air, can bring the air to saturation at the same temperature adiabatically. An adiabatic saturator is a device using which one can measure theoretically the adiabatic saturation temperature of air.

As shown in Fig.2.1 an adiabatic saturator is a device in which air flows through an infinitely long duct containing water. As the air comes in contact with water in the duct, there will be heat and mass transfer between water and air. If the duct is infinitely long, then at the exit, there would exist perfect equilibrium between air and water at steady state. Air at the exit would be fully saturated and its temperature is equal to that of water temperature. The device is adiabatic as the walls of the chamber are thermally insulated. In order to continue the process, make-up water has to be provided to compensate for the amount of water evaporated into the air. The temperature of the make-up water is controlled so that it is the same as that in the duct.

After the adiabatic saturator has achieved a steady-state condition, the temperature indicated by the thermometer immersed in the water is the *thermodynamic wet-bulb temperature*. The thermodynamic wet bulb temperature will be less than the entering air DBT but greater than the dew point temperature.

Certain combinations of air conditions will result in a given sump temperature, and this can be defined by writing the energy balance equation for the adiabatic saturator. Based on a unit mass flow rate of dry air, this is given by:

$$h_1 = h_2 - (W_2 - W_1)h_f \quad \dots\dots\dots (2.21)$$

where h_f is the enthalpy of saturated liquid at the sump or thermodynamic wet-bulb temperature, h_1 and h_2 are the enthalpies of air at the inlet and exit of the adiabatic saturator, and W_1 and W_2 are the humidity ratio of air at the inlet and exit of the adiabatic saturator, respectively.

It is to be observed that the thermodynamic wet-bulb temperature is a thermodynamic property, and is independent of the path taken by air. Assuming the humid specific heat to be constant, from the enthalpy balance, the thermodynamic wet-bulb temperature can be written as:

$$t_2 = t_1 - \frac{h_{fg,2}}{c_{pm}} (w_2 - w_1) \quad \dots\dots\dots (2.22)$$

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where $h_{fg,2}$ is the latent heat of vaporization at the saturated condition 2. Thus measuring the dry bulb (t_1) and wet bulb temperature (t_2) one can find the inlet humidity ratio (W_1) from the above expression as the outlet saturated humidity ratio (W_2) and latent heat of vaporizations are functions of t_2 alone (at fixed barometric pressure).

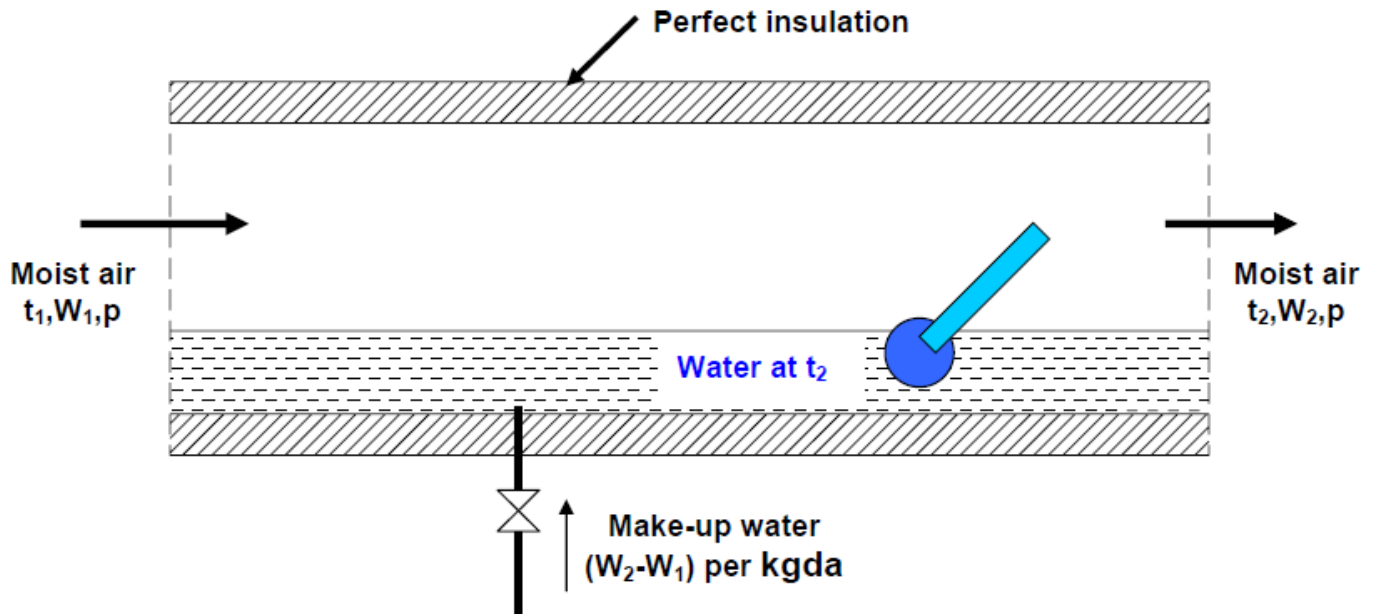


Fig.2.1 The process of adiabatic saturation of air

Example:

Moist air at 26°C dry bulb and 16°C wet bulb enters to adiabatic device. The air leaving the device completely saturated at standard atmospheric pressure ($P_{atm}=101.325$ kPa)

Water flow at 16°C, Calculate:

- moisture content at the entry and exit W_1 & W_2
- degree of saturation at the entry μ_1
- specific volume at the entry and exit v_1 & v_2
- enthalpy at the entry and exit h_1 & h_2

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Solution:

a) From Table A-2 at $t_w = 16^\circ\text{C}$, $P_{wss} = 1.818 \text{ kPa}$

$$P_v = 1.818 - 101.325 \times 6.66 \times 10^{-4} (26 - 16) = 1.1432 \text{ kPa}$$

$$W_1 = 0.622(1.1432)/(101.325 - 1.1432) = 0.00709 \text{ kg/kg dry air}$$

$$W_2 = W_{sw} = 0.622(1.818)/(101.325 - 1.818) = 0.01136 \text{ kg/kg dry air}$$

b) $\mu = [W/W_{ss1}]_{t,p}$

where (W_{ss}) is the moisture content for saturated air at entry state

From Table A-2 at $t = 26^\circ\text{C}$, $P_{ss} = 3.360 \text{ kPa}$

$$W_{ss1} = 0.622(3.360)/(101.325 - 3.360) = 0.0213 \text{ kg/kg dry air}$$

Then; $\mu = (0.00709)/(0.0213) = 0.333 = 33.3\%$

c) $P_{a1} = P_B - P_1 = 101.325 - 1.1432 = 100.182 \text{ kPa}$

$$P_{a2} = P_B - P_2 = 101.325 - 1.818 = 99.507 \text{ kPa}$$

$$v_{a1} = R_a T_{a1} / P_{a1} = 0.287 (273 + 26) / (100.183) = 0.857 \text{ m}^3/\text{kg dry air}$$

$$v_{a2} = R_a T_{a2} / P_{a2} = 0.287 (273 + 16) / (99.507) = 0.833 \text{ m}^3/\text{kg dry air}$$

d) $h = (1.007 t_d - 0.026) + W(2501 + 1.84 t_d)$

$$h_1 = (1.007 \times 26 - 0.026) + 0.00709(2501 + 1.84 \times 26) = 44.23 \text{ kJ/kg dry air}$$

$$h_2 = (1.007 \times 16 - 0.026) + 0.01136(2501 + 1.84 \times 16) = 44.83 \text{ kJ/kg dry air}$$

Also, can be calculate (h_2) from Eqn.(2.21)

$$h_2 = h_1 + (W_2 - W_1) h_f$$

From Table A-2 at $t = 16^\circ\text{C}$ $h_f = 67.19 \text{ kJ/kg}$

$$h_2 = 44.23 + (0.01136 - 0.00709) \times 67.19 = 44.517 \text{ kJ/kg dry air}$$

- **Note here the enthalpy of the dry air decreases about 10 kJ/kg (from 26.208 to 16.138) while the enthalpy of water vapor is increases within the same amount 10 kJ/kg (from 18.071 to 28.745). This is an important fact in adiabatic saturation process.**

Chapter Two

Exercises

1. The atmospheric condition of air are 25°C dry bulb temperature and moisture content of 0.01 kg/kg dry air. Find:

(a) partial pressure of vapor (b) relative humidity (c) dew point temperature.

[Ans. 0.016 bar , 50.6% , 14.1°C]

2. A sling psychrometer reads 40°C dry bulb temperature and 28°C wet bulb temperature. Calculate the following:

(a) moisture content (b) relative humidity (c) vapor density in air

(d) dew point temperature (e) enthalpy of mixture per kg of dry air.

[Ans. 0.019 kg/kg of dry air, 40.7% , 0.0208 kg/m^3 , 24°C , 88.38 kJ/kg dry air]

3. A sample of moist air has a dry bulb temperature of 25°C and a relative humidity of 50% . The barometric pressure is 740 mm of Hg. Calculate:

(a) partial pressure of water vapor and dry air (b) dew point temperature

(c) specific humidity (moisture content) (d) enthalpy of air

[Ans. 0.01583 bar , 14°C , 0.0101 kg/kg dry air, 50.81 kJ/kg dry air]

4. The moist air exists at a total pressure of 1.01325 bar and 25°C dry bulb temperature. If the degree of saturation is 50% , determine the following using steam tables:

(a) moisture content (b) dew point temperature (c) specific volume of moist air.

[Ans. 10.03 g/kg dry air, 14°C , $0.857\text{ m}^3/\text{kg}$]

5. The atmospheric conditions of air are 35°C dry bulb temperature, 60% relative humidity and 1.01325 bar pressure. If 0.005 kg of moisture per kg of dry air is removed, the temperature becomes 25°C . Determine the final relative humidity and dew point temperature.

[Ans. 88.6% , 23°C]

6. An atmospheric air enters the adiabatic saturator at 33°C dry bulb temperature and 23°C wet bulb temperature. The barometric pressure is 740 mm Hg. Determine the moisture content and vapor pressure at 33°C .

[Ans. 0.012 kg/kg dry air, 13 mm Hg]

Chapter Two

7. Air at 40°C DBT and 15% RH is passed through the adiabatic humidifier at the rate of 200 m³/min. The outlet conditions of air are 25°C DBT and 20°C WBT. Find:

- (a) dew point temperature (b) relative humidity of exit air
(c) amount of water vapor added to the air per minute.

[Ans. 17.8°C, 65%, 1.26 kg/min]

8. The atmospheric air has 35°C DBT and 50% RH and standard atmospheric pressure. Find:

- (a) wet bulb temperature (b) moisture content (c) dew point temperature (d) enthalpy

[Ans. 26.20°C, 0.0178 kg/kg_{d.a}, 23°C, 81 kJ/kg]

9. An atmospheric air at 15°C DBT and 80% RH is supplied to the heating chamber at the rate of 100 m³/min.. The leaving air has a temperature of 22°C without change in its moisture content. Determine the heat added to the air per min. and final relative humidity of the air.

[Ans. 865.3 kJ/min, 53%]

10. Moist air at 20°C DBT and the barometric pressure is 82.5 kPa. Calculate the enthalpy and the degree of saturation, if the air was:

- (a) saturated (b) at a relative humidity of 50%

[Ans. 66.124 kJ/kg, 1.0 and 42.792 kJ/kg, 0.493]

Chapter Three

The Psychrometry of Air Conditioning Processes

3.1 The Psychrometric Chart

This provides a picture of the way in which the state of moist air alters as an air conditioning process takes place or a physical change occurs. Familiarity with the psychrometric chart is essential for a proper understanding of air conditioning.

Any point on the chart is termed a *state point*, the location of which, at a given barometric pressure, is fixed by any two of the psychrometric properties discussed in chapter 2. It is customary and convenient to design charts at a constant barometric pressure because barometric pressure does not alter greatly over much of the inhabited surface of the earth. When the barometric pressure is significantly different from the standard adopted for the chart or psychrometric tables to hand, then the required properties can be calculated using the equations derived earlier.

The British standard is that adopted by the Chartered Institution of Building Services Engineers for their Tables of Psychrometric Data and for their psychrometric chart. It is 101.325 kPa. The American standard is also 101.325 kPa and this value is used by the American Society of Heating, Refrigeration and Air Conditioning Engineers.

The psychrometric chart published by the CIBSE uses two fundamental properties, mass and energy, in the form of moisture content and enthalpy, as co-ordinates. As a result, mixture states lie on the straight line which joins the state points of the two constituents. Lines of constant dry-bulb temperature are virtually straight but divergent, only the isotherm for 30°C being vertical. The reason for this is that to preserve the usual appearance of a psychrometric chart, in spite of choosing the two fundamental properties as co-ordinates, the co-ordinate axes are oblique, not rectangular. Hence, lines of constant enthalpy are both straight and parallel, as are lines of constant moisture content. Since both these properties are taken as linear, the lines of constant enthalpy are equally spaced as are, also, the lines of constant moisture content. This is not true of the lines of constant humid

volume and constant wet-bulb temperature, which are slightly curved and divergent. Since their curvature is only slight in the region of practical use on the chart, they can be regarded as straight without significant error resulting. In the sketches of psychrometric charts used throughout this text to illustrate changes of state, only lines of percentage saturation are shown curved. All others are shown straight, and dry-bulb isotherms are shown as vertical, for convenience.

The chart also has a protractor which allows the value of the ratio of the sensible heat gain to the total heat gain to be plotted on the chart. This ratio is an indication of the slope

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of the room ratio line and is of value in determining the correct supply state of the air that must be delivered to a conditioned space. The zero value for the ratio is parallel to the isotherm for 30°C because the enthalpy of the added vapour depends on the temperature at which evaporation takes place, it being assumed that most of the latent heat gain to the air in a conditioned room is by evaporation from the skin of the occupants and that their skin surface temperature is about 30°C.

Figure 3.1 shows a state point on a psychrometric chart.

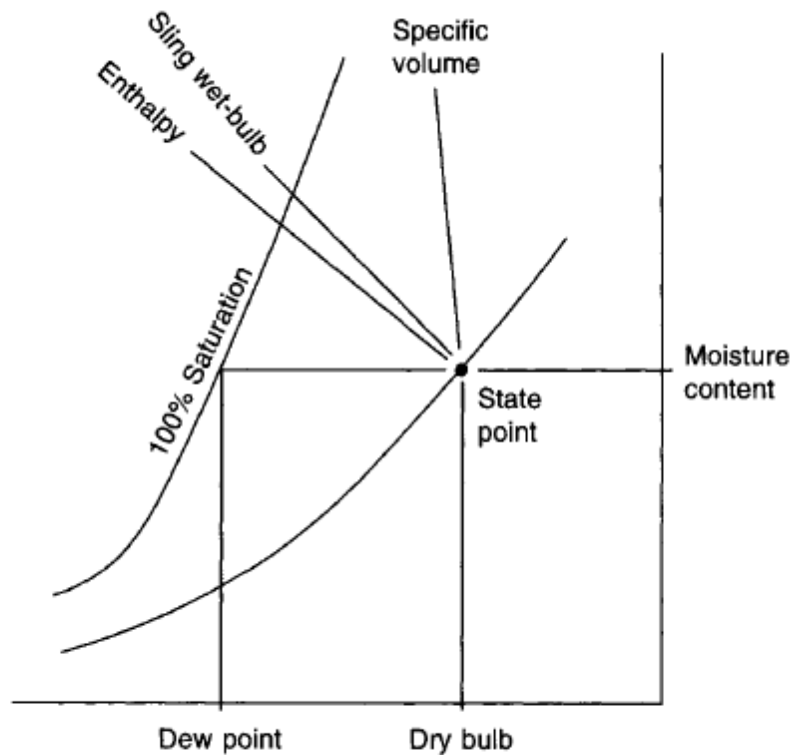


Fig. 3.1 The variables shown on the psychrometric chart.

3.2 Mixtures

Figure 3.2 shows what happens when two airstreams meet and mix adiabatically. Moist air at state 1 mixes with moist air at state 2, forming a mixture at state 3. The principle of the conservation of mass allows two mass balance equations to be written:

$$m_{a1} + m_{a2} = m_{a3} \text{ for the dry air and}$$

$$g_1 m_{a1} + g_2 m_{a2} = g_3 m_{a3} \text{ for the associated water vapour}$$

Hence

$$(g_1 - g_3)m_{a1} = (g_3 - g_2)m_{a2}$$

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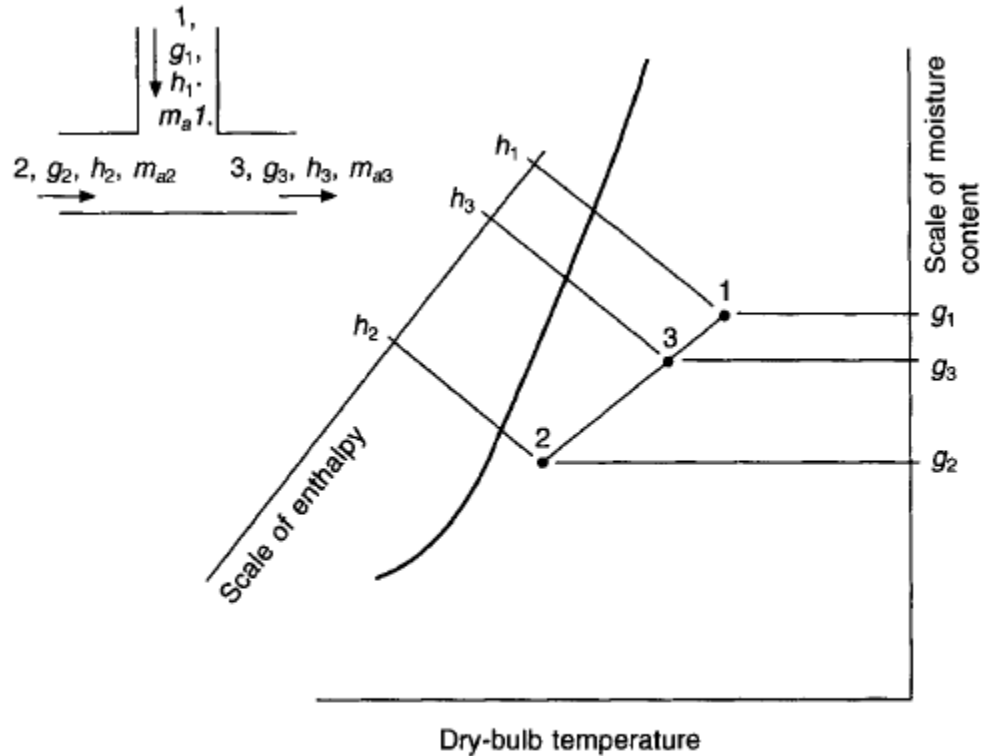


Fig. 3.2 The adiabatic mixing of two airstreams.

Therefore

$$\frac{g_1 - g_3}{g_3 - g_2} = \frac{m_{a2}}{m_{a1}}$$

Similarly, making use of the principle of the conservation of energy,

$$\frac{h_1 - h_3}{h_3 - h_2} = \frac{m_{a2}}{m_{a1}}$$

From this it follows that the three state points must lie on a straight line in a mass-energy co-ordinate system. When two airstreams mix adiabatically, the mixture state lies on the straight line which joins the state points of the constituents, and the position of the mixture state point is such that the line is divided inversely as the ratio of the masses of dry air in the constituent airstreams.

EXAMPLE 3.1

Moist air at a state of 60°C dry-bulb, 32.1°C wet-bulb (sling) and 101.325 kPa barometric pressure mixes adiabatically with moist air at 5°C dry-bulb, 0.5°C wet-bulb (sling) and 101.325 kPa barometric pressure. If the masses of dry air are 3 kg and 2 kg, respectively, calculate the moisture content, enthalpy and dry-bulb temperature of the mixture.

Answer

$$g_1 = 18.400 \text{ g per kg dry air}$$

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$$g_2 = 2.061 \text{ g per kg dry air}$$

$$h_1 = 108.40 \text{ kJ per kg dry air}$$

$$h_2 = 10.20 \text{ kJ per kg dry air}$$

The principle of the conservation of mass demands that

$$\begin{aligned} g_1 m_{a1} + g_2 m_{a2} &= g_3 m_{a3} \\ &= g_3 (m_{a1} + m_{a2}) \end{aligned}$$

hence

$$\begin{aligned} g_3 &= \frac{g_1 m_{a1} + g_2 m_{a2}}{m_{a1} + m_{a2}} \\ &= \frac{18.4 \times 3 + 2.061 \times 2}{3 + 2} \\ &= \frac{59.322}{5} \\ &= 11.864 \text{ g per kg dry air} \end{aligned}$$

Similarly, by the principle of the conservation of energy,

$$\begin{aligned} h_3 &= \frac{h_1 m_{a1} + h_2 m_{a2}}{m_{a1} + m_{a2}} \\ &= \frac{108.40 \times 3 + 10.20 \times 2}{3 + 2} \\ &= 69.12 \text{ kJ per kg dry air} \end{aligned}$$

To determine the dry-bulb temperature, the following practical equation must be used

$$h = (1.007t - 0.026) + g(2501 + 1.84t)$$

Substituting the values calculated for moisture content and enthalpy, this equation can be solved for temperature:

$$h = 69.12 = (1.007t - 0.026) + 0.01186(2501 + 1.84t)$$

$$t = \frac{39.48}{1.029} = 38.4^\circ\text{C}$$

On the other hand, if the temperature were calculated by proportion, according to the masses of the dry air in the two mixing airstreams, a slightly different answer results:

$$\begin{aligned} t &= \frac{3 \times 60 + 2 \times 5}{5} \\ &= 38^\circ\text{C} \end{aligned}$$

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EXAMPLE 3.2

A stream of moist air at a state of 21°C dry-bulb and 14.5°C wet-bulb (sling) mixes with another stream of moist air at a state of 28°C dry-bulb and 20.2°C wet-bulb (sling), the respective masses of the associated dry air being 3 kg and 1 kg. With the aid of CIBSE tables of psychrometric data calculate the dry-bulb temperature of the mixture (a) using the principles of conservation of energy and of mass and, (b), using a direct proportionality between temperature and mass.

3.3 Sensible heating and cooling

Sensible heat transfer occurs when moist air flows across a heater battery or over the coils of a sensible cooler. In the heater, the temperature of the medium used to provide the heat is not critical. The sole requirement for heat transfer is that the temperature shall exceed the final air temperature. In sensible cooling there is a further restriction: the lowest water temperature must not be so low that moisture starts to condense on the cooler coils. If such condensation does occur, through a poor choice of chilled water temperature, then the process will no longer be one of sensible cooling since dehumidification will also be taking place.

Figure 3.3 shows the changes of state which occur, sketched upon a psychrometric chart. The essence of both processes is that the change of state must occur along a line of constant moisture content. The variations in the physical properties of the moist air, for the two cases, are summarised below:

	<i>Sensible heating</i>	<i>Sensible cooling</i>
Dry-bulb	increases	decreases
Enthalpy	increases	decreases
Humid volume	increases	decreases
Wet-bulb	increases	decreases
Percentage saturation	decreases	increases
Moisture content	constant	constant
Dew point	constant	constant
Vapour pressure	constant	constant

EXAMPLE 3.3

Calculate the load on a battery which heats $1.5 \text{ m}^3 \text{ s}^{-1}$ of moist air, initially at a state of 21°C dry-bulb, 15°C wet-bulb (sling) and 101.325 kPa barometric pressure, by 20 degrees. If low temperature hot water at 85°C flow and 75°C return is used to achieve this, calculate the flow rate necessary, in kilograms of water per second.

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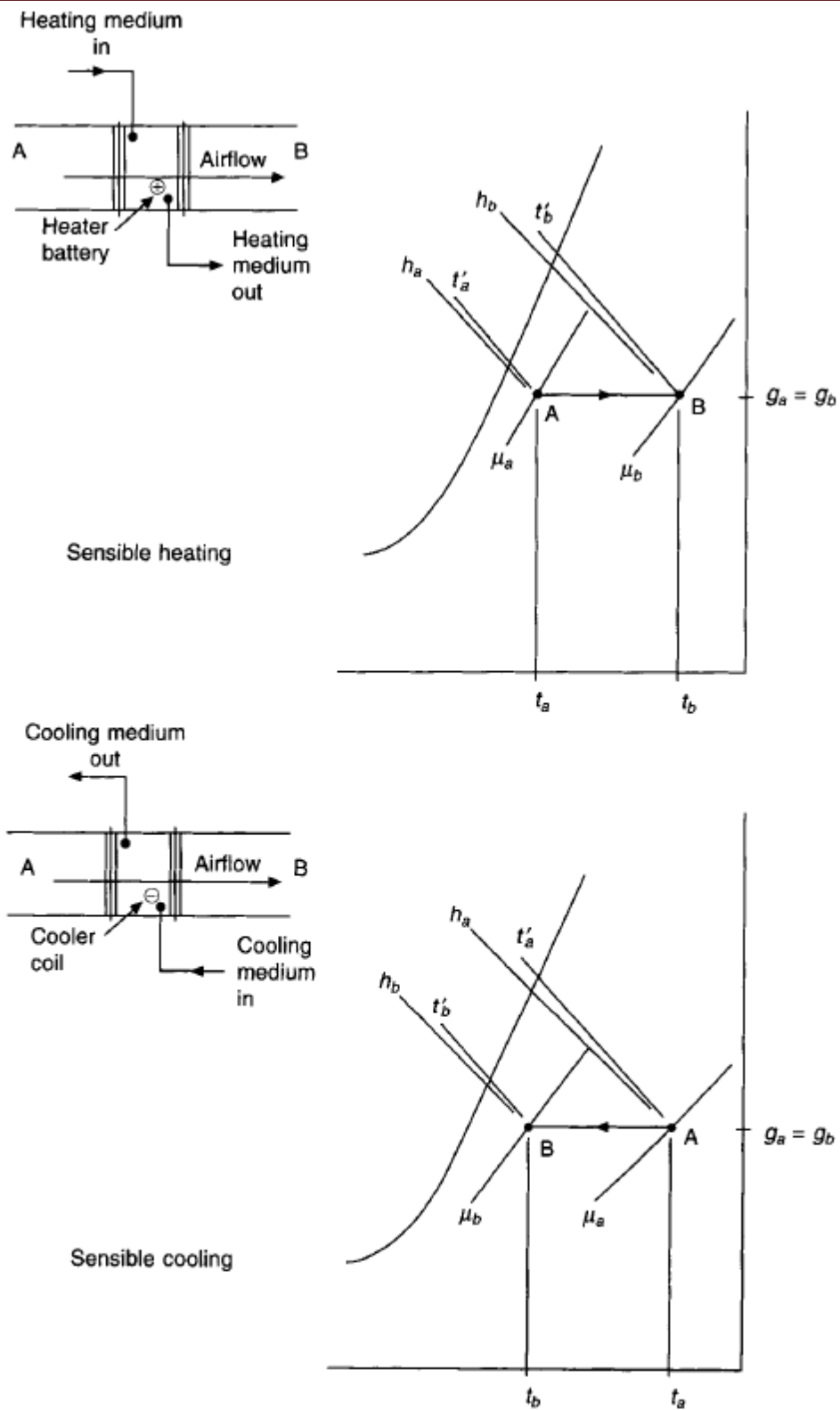


Fig. 3.3 Psychrometry for sensible heating and cooling.

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Answer

$$\text{Heating load} = \left(\frac{\text{mass flow of moist air expressed}}{\text{in kg s}^{-1} \text{ of associated dry air}} \right) \times \left(\frac{\text{increase in enthalpy of moist air expressed}}{\text{in kJ per kg of associated dry air}} \right)$$

From CIBSE tables of psychrometric data (or, less accurately, from the CIBSE psychrometric chart), the initial enthalpy is found to be 41.88 kJ kg⁻¹, the moisture content to be 8.171 g kg⁻¹ and the humid volume to be 0.8439 m³ per kg of dry air. Since the air is being heated by 20 degrees, reference must now be made to tables in order to determine the enthalpy at the same moisture content as the initial state but at a dry-bulb temperature of 41°C. By interpolation, the enthalpy of the moist air leaving the heater battery is found to be 62.31 kJ per kg of dry air.

$$\text{Heating load} = \left(\frac{1.5}{0.8439} \right) \times (62.31 - 41.88) = 36.3 \text{ kW}$$

$$\text{Flow rate of LTHW} = \frac{36.3}{(85^\circ - 75^\circ) \times 4.2} = 0.864 \text{ kg s}^{-1}$$

where 4.2 kJ/kg K is the specific heat capacity of water.

EXAMPLE 3.4

Calculate the load on a cooler coil which cools the moist air mentioned in example 3.3 by 5 degrees. What is the flow rate of chilled water necessary to effect this cooling if flow and return temperatures of 10°C and 15°C are satisfactory?

Answer

The initial enthalpy and humid volume are the same as in the first example. The final dry-bulb temperature of the moist air is 16°C but its moisture content is still 8.171 g per kg of dry air. At this state, its enthalpy is found from tables to be 36.77 kJ per kg of dry air.

$$\text{Cooling load} = \left(\frac{1.5}{0.8439} \right) \times (41.88 - 36.77) = 9.1 \text{ kW}$$

$$\text{Chilled water flow rate} = \frac{9.1}{5 \times 4.2} = 0.433 \text{ kg s}^{-1}$$

It is to be noted that, for sensible cooling, the selection of cooler coils and the choice of chilled water flow temperature require some care. See section 10.6.

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3.4 Dehumidification

There are four principal methods whereby moist air can be dehumidified:

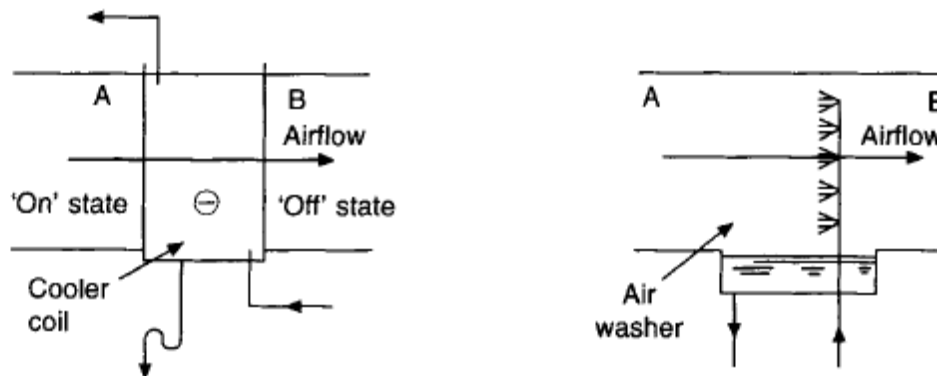
- (i) cooling to a temperature below the dew point,
- (ii) adsorption,
- (iii) absorption,
- (iv) compression followed by cooling.

The first method forms the subject matter of this section.

Cooling to a temperature below the dew point is done by passing the moist air over a cooler coil or through an air washer provided with chilled water.

Figure 3.4 shows on a sketch of a psychrometric chart what happens when moist air is cooled and dehumidified in this fashion. Since dehumidification is the aim, some of the spray water or some part of the cooler coil must be at a temperature less than the dew point of the air entering the equipment. In the figure, t_d is the dew point of the moist air 'on' the coil or washer. The temperature t_c , corresponding to the point C on the saturation curve, is termed the *apparatus dew point*. This term is in use for both coils and washers but, in the case of cooler coils alone, t_c is also sometimes termed the *mean coil surface temperature*. The justification for this latter terminology is offered in chapter 10.

For purposes of carrying out air conditioning calculations, it is sufficient to know the



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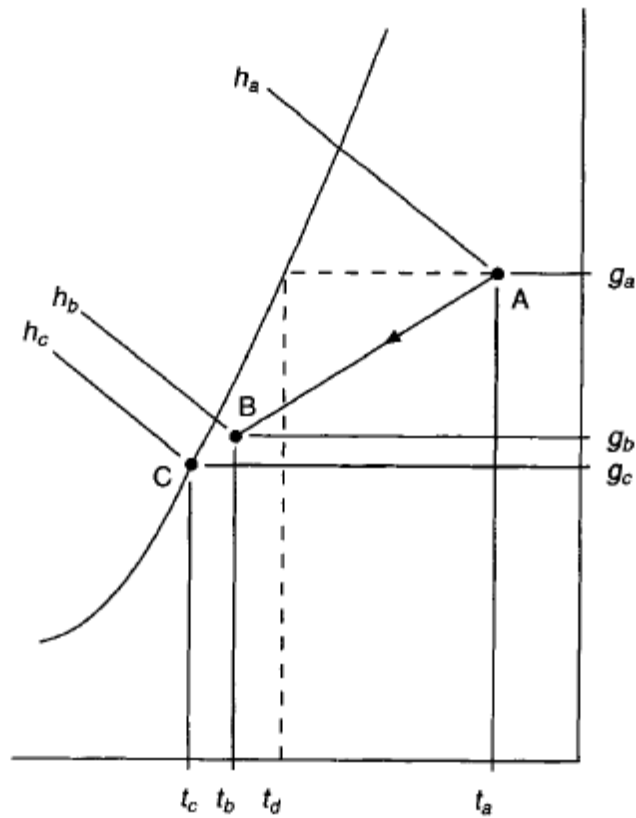


Fig. 3.4 Cooling and dehumidification by a cooler coil or an air washer.

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state A of the moist air entering the coil, the state B of the air leaving the coil, and the mass flow of the associated dry air. What happens to the state of the air as it passes between points A and B is seldom of more than academic interest. Consequently, it is quite usual to show the change of state between the 'on' and the 'off' conditions as occurring along a straight line. In fact, as will be seen in chapter 10, the change of state follows a curved path, the curvature of which is a consequence of the heat transfer characteristics of the process, not of the construction of the psychrometric chart.

It can be seen from Figure 3.4 that the moisture content of the air is reduced, as also is its enthalpy and dry-bulb temperature. The percentage saturation, of course, increases. It might be thought that the increase of humidity would be such that the 'off' state, represented by the point B, would lie on the saturation curve. This is not so for the very good reason that no air washer or cooler coil is a hundred per cent efficient. It is unusual to speak of the efficiency of a cooler coil. Instead, the alternative terms, *contact factor* and *by-pass factor*, are used. They are complementary values and contact factor, sometimes denoted by β , is defined as

$$\begin{aligned}\beta &= \frac{g_a - g_b}{g_a - g_c} \\ &= \frac{h_a - h_b}{h_a - h_c}\end{aligned}\quad (3.1)$$

Similarly, by-pass factor is defined as

$$\begin{aligned}(1 - \beta) &= \frac{g_b - g_c}{g_a - g_c} \\ &= \frac{h_b - h_c}{h_a - h_c}\end{aligned}\quad (3.2)$$

It is sufficient, for all practical purposes, to assume that both these expressions can be rewritten in terms of dry-bulb temperature, namely that

$$\beta = \frac{t_a - t_b}{t_a - t_c}\quad (3.3)$$

and that

$$(1 - \beta) = \frac{t_b - t_c}{t_a - t_c}\quad (3.4)$$

It is much less true to assume that they can also be written, without sensible error, in terms of wet-bulb temperature, since the scale of wet-bulb values on the psychrometric chart is not at all linear. The assumption is, however, sometimes made for convenience, provided the values involved are not very far apart and that some inaccuracy can be tolerated in the answer.

Typical values of β are 0.82 to 0.92 for practical coil selection in the UK. In hot, humid climates more heat transfer surface is necessary and higher contact factors are common.

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EXAMPLE 3.5

$1.5 \text{ m}^3 \text{ s}^{-1}$ of moist air at a state of 28°C dry-bulb, 20.6°C wet-bulb (sling) and 101.325 kPa flows across a cooler coil and leaves the coil at 12.5°C dry-bulb and 8.336 g per kg of dry air.

Determine (a) the apparatus dew point, (b) the contact factor and (c) the cooling load.

Answer

Figure 3.5 shows the psychrometric changes involved and the values immediately known from the data in the question. From CIBSE tables (or from a psychrometric chart) it is established that $h_a = 59.06 \text{ kJ kg}^{-1}$ and that $h_b = 33.61 \text{ kJ kg}^{-1}$. The tables also give a value for the humid volume at the entry state to the coil, of $0.8693 \text{ m}^3 \text{ kg}^{-1}$.

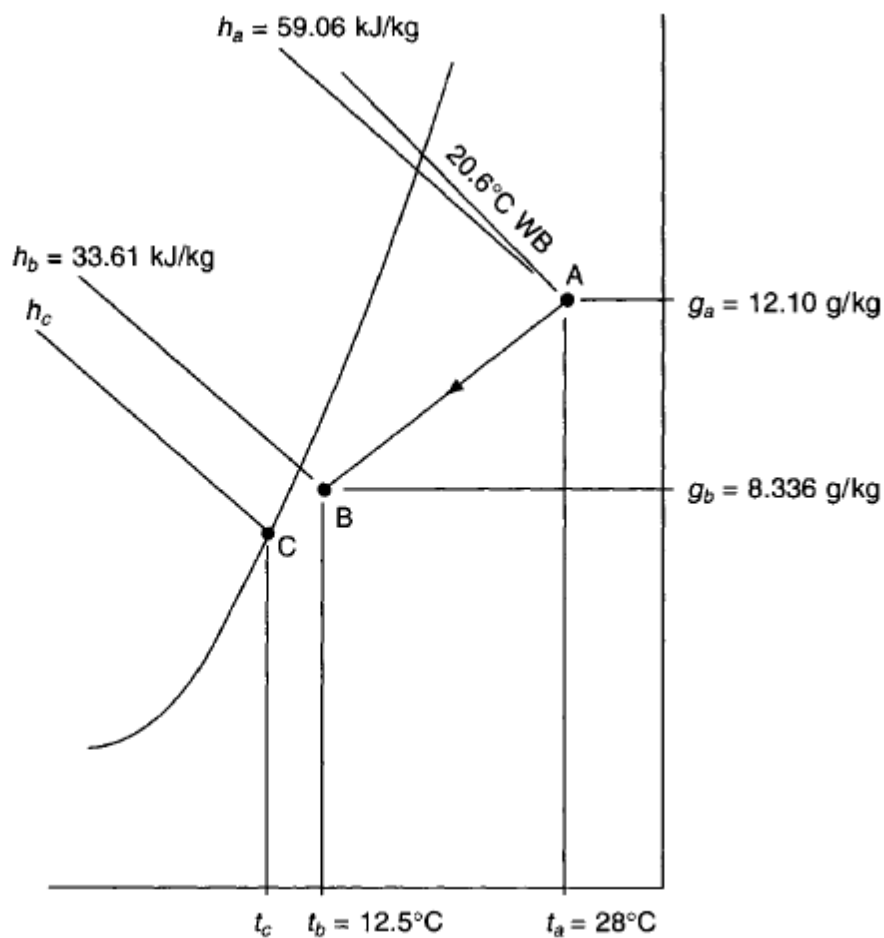


Fig. 3.5 The psychrometry for Example 3.5.

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(a) Mark state points A and B on a psychrometric chart. Join them by a straight line and extend the line to cut the saturation curve. Observation of the point of intersection shows that this is at a temperature of 10.25°C. It is not easy to decide this value with any exactness unless a large psychrometric chart is used. However, it is sufficiently accurate for present purposes (and for most practical purposes) to take the value read from an ordinary chart. Hence the apparatus dew point is 10.25°C.

(b) Either from the chart or from tables, it can be established that the enthalpy at the apparatus dew point is 29.94 kJ kg⁻¹. Using the definition of contact factor,

$$\beta = \frac{59.06 - 33.61}{59.06 - 29.94} = 0.874$$

One might also determine this from the temperatures:

$$= \frac{28^\circ - 12.5^\circ}{28^\circ - 10.25^\circ} = 0.873$$

Clearly, in view of the error in reading the apparatus dew point from the chart, the value of β obtained from the temperatures is quite accurate enough in this example and, in fact, in most other cases.

(c) Cooling load = mass flow rate \times decrease of enthalpy

$$= \frac{1.5}{0.8693} \times (59.06 - 33.61) = 43.9 \text{ kW}$$

3.5 Humidification

This, as its name implies, means that the moisture content of the air is increased. This may be accomplished by either water or steam but the present section is devoted to the use of water only, section 3.7 being used for steam injection into moving airstreams.

It is customary to speak of the humidifying efficiency or the effectiveness of an air washer (although neither term is universally accepted) rather than a contact or by-pass factor. There are several definitions, some based on the extent to which the dry-bulb temperature of the entering moist airstream approaches its initial wet-bulb value, and others based on the change of state undergone by the air. In view of the fact that the psychrometric chart currently in use by the Chartered Institution of Building Services

Engineers is constructed with mass (moisture content) and energy (enthalpy) as oblique, linear co-ordinates, the most suitable definition to use with the chart is that couched in terms of these fundamentals. There is the further advantage that such a definition of effectiveness, E , is identical with the definition of contact factor, β , used for cooler coils.

Although humidifying efficiency is often expressed in terms of a process of adiabatic saturation, this is really a special case, and is so regarded here.

Figure 3.6(a) shows an illustration of the change of state experienced by an airstream as it passes through a spray chamber.

The effectiveness of the spray chamber is then defined by

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$$E = \frac{h_b - h_a}{h_c - h_a} \quad (3.5)$$

and humidifying efficiency is defined by

$$\eta = 100E \quad (3.6)$$

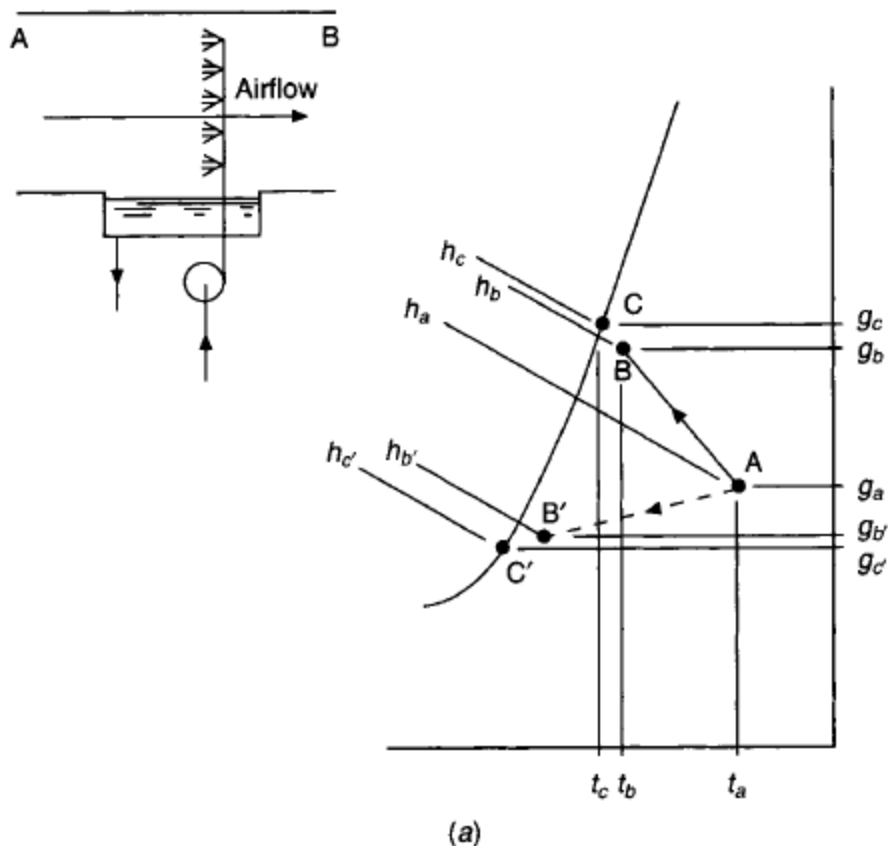
It is evident that, because moisture content is the other linear co-ordinate of the psychrometric chart and the points A, B and C lie on the same straight line, C being obtained by the extension of the line joining AB to cut the saturation curve, then it is possible to put forward an alternative and equally valid definition of effectiveness, expressed in terms of moisture content:

$$E = \frac{g_b - g_a}{g_c - g_a} \quad (3.7)$$

and

$$\eta = 100E, \text{ as before}$$

Figure 3.6(a) shows that h_b is greater than h_a .



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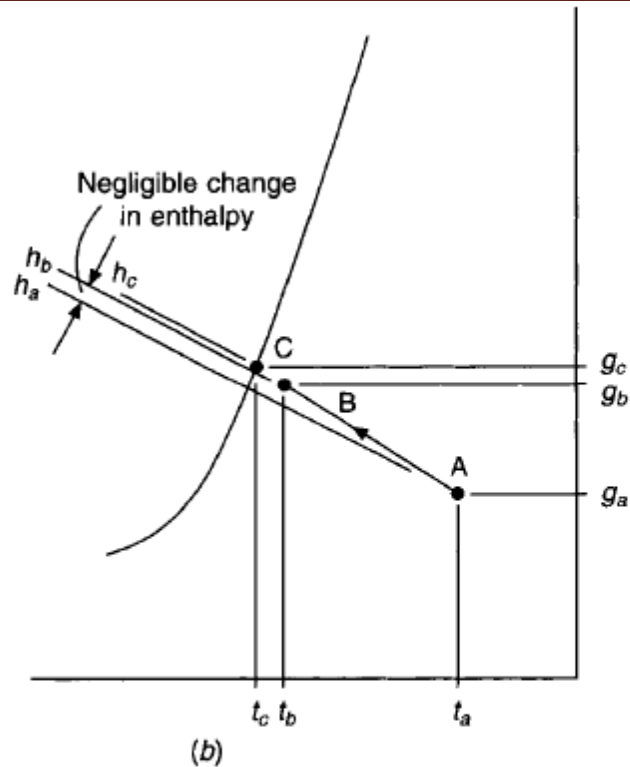


Fig. 3.6 (a) The change of state when air flows through a spray chamber having spray water at a controlled temperature. (b) Adiabatic saturation when the spray water is entirely recirculated and neither cooled nor heated.

If the spray water had been chilled, instead of heated, then a change of state might have been as shown dotted on Figure 3.6(a), from the point A to the point B'. Under these circumstances, the effectiveness would have been expressed by

$$E = \frac{h_a - h_{b'}}{h_a - h_{c'}}$$

or

$$E = \frac{g_a - g_{b'}}{g_a - g_{c'}}$$

Consider the special case of adiabatic saturation: for this to occur it is necessary that

- (i) the spray water is totally recirculated, no heat exchanger being present in the pipelines or in the washer tank;
- (ii) the spray chamber, tank and pipelines are perfectly lagged; and
- (iii) the feed water supplied to the system to make good the losses due to evaporation is at the temperature of adiabatic saturation

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Under these conditions it may be assumed that the change of state follows a line of constant wet-bulb temperature (since the Lewis number for air–water vapour mixtures is unity), the system having been given sufficient time to settle down to steady-state operation. There must be a change of enthalpy during the process because feed water is being supplied at the wet-bulb temperature (virtually) and this will not, as a general rule, equal the datum temperature for the enthalpy of water, 0°C. Strictly speaking then, it is incorrect to speak of the process as being an adiabatic one. The use of the term stems from the phrase ‘temperature of adiabatic saturation’ and so its use is condoned. One thing is fairly certain though: at the temperatures normally encountered the change of enthalpy during the process is negligible, and so effectiveness must be expressed in terms of change of moisture content. Figure 3.6(b) shows a case of adiabatic saturation, in which it can be seen that there is no significant alteration in enthalpy although there is a clear change in moisture content. It can also be seen that a fall in temperature accompanies the rise in moisture content. This temperature change provides an approximate definition of effectiveness or efficiency which is most useful, and, for the majority of practical applications, sufficiently accurate. It is

$$E = \frac{t_a - t_b}{t_a - t_c} \quad (3.8)$$

The way in which the psychrometric chart is constructed precludes the possibility of this being an accurate expression; lines of constant dry-bulb temperature are not parallel and equally spaced, such properties being exclusive to enthalpy and moisture content.

EXAMPLE 3.6

1.5 m³ s⁻¹ of moist air at a state of 15°C dry-bulb, 10°C wet-bulb (sling) and 101.325 kPa barometric pressure, enters the spray chamber of an air washer. The humidifying efficiency of the washer is 90 per cent, all the spray water is recirculated, the spray chamber and the tank are perfectly lagged, and mains water at 10°C is supplied to make good the losses due to evaporation.

Calculate (a) the state of the air leaving the washer, (b) the rate of flow of make-up water from the mains.

Answer

(a) This is illustrated in Figure 3.7.

Using the definitions of humidifying efficiency quoted in equations (3.6) and (3.7), and referring to tables of psychrometric data for the properties of moist air at states A and C, one can write

$$\frac{90}{100} = \frac{g_b - 5.558}{7.659 - 5.558}$$

Hence,

$$g_b = 7.449 \text{ g per kg of dry air}$$

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The state of the moist air leaving the air washer is 10°C wet-bulb (sling), 7.449 g per kg and 101.325 kPa barometric pressure. The use of equation (3.8) shows that the approximate dry-bulb temperature at exit from the washer is 10.5°C.

(b) The amount of water supplied to the washer must equal the amount evaporated. From tables (or less accurately from a psychrometric chart), the humid volume at state A is 0.8232 m³ kg⁻¹. Each kilogram of dry air passing through the spray chamber has its associated water vapour augmented by an amount equal to $g_b - g_a$, that is, by 1.891 g.

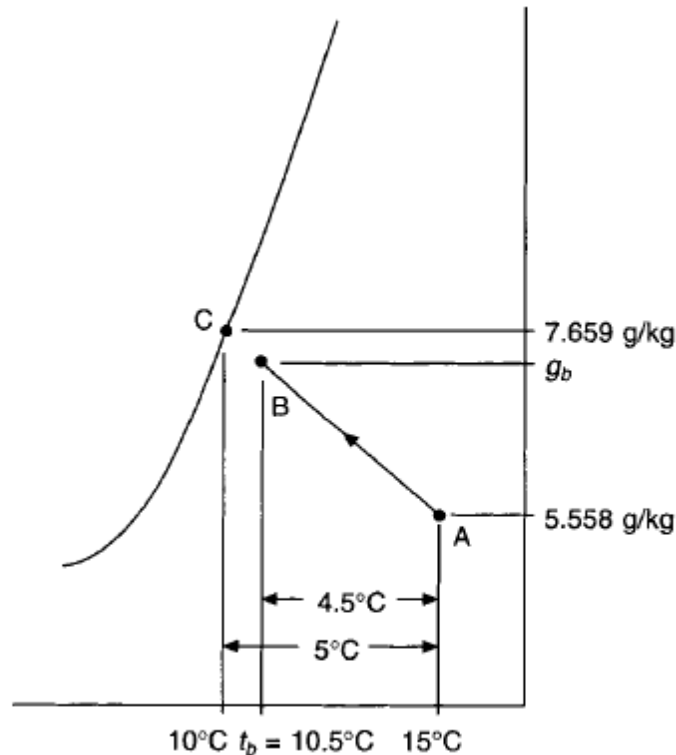


Fig. 3.7 Adiabatic saturation along a wet-bulb line.

Thus, the rate of make-up is

$$\begin{aligned}
 &= \frac{1.5 \times 0.001891 \times 3600}{0.8232} \\
 &= 12.40 \text{ kg of water per hour}
 \end{aligned}$$

3.6 Water injection

The simplest case to consider, and the one that provides the most insight into the change of state of the airstream subjected to humidification by the injection of water, is where all the injected water is evaporated. Figure 3.8 shows what happens when total evaporation occurs.

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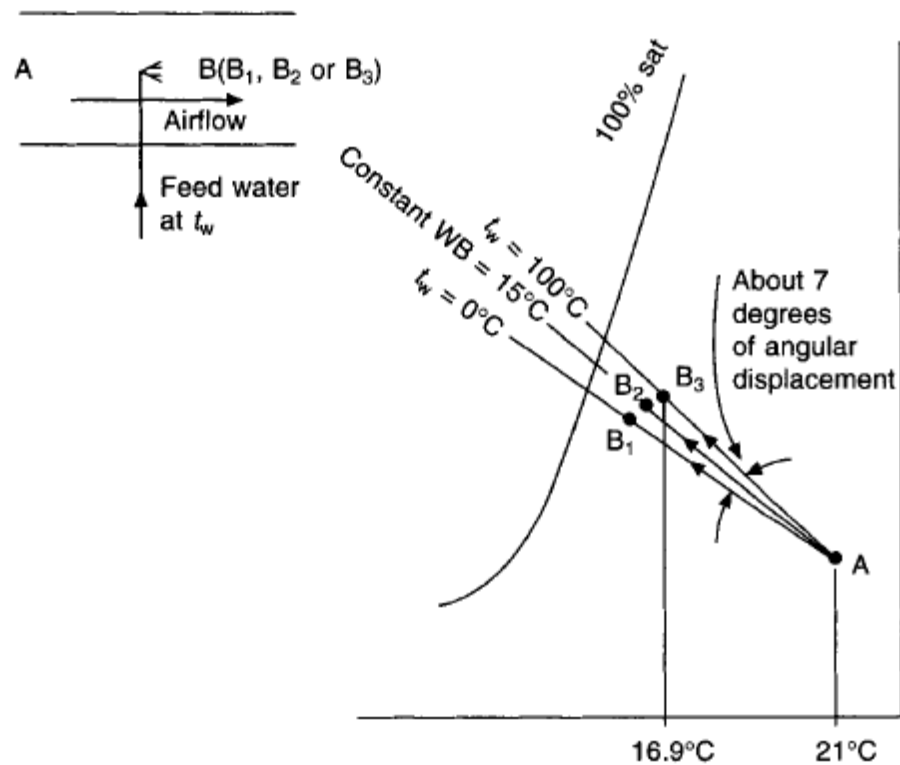


Fig. 3.8 A humidification process by the entire evaporation of injected spray water.

Air enters a spray chamber at state A and leaves it at state B, all the injected water being evaporated, none falling to the bottom of the chamber to run to waste or to be recirculated. The feed-water temperature is t_w . It is important to realise that since total evaporation has occurred, state B must lie nearer to the saturation curve, but just how much nearer will depend on the amount of water injected.

Two equations, a heat balance and a mass balance, provide the answer required.

Striking a heat balance, we can write

$$h_a + h_w = h_b$$

and, in a similar way, the mass balance may be written as

$$m_a + m_w = m_b$$

Knowing the amount of feed water evaporated, the mass balance gives the necessary information about the moisture content of the airstream leaving the area of the water injection. Applying the mass balance to the water vapour only (since the associated kilogram of dry air may be ignored),

$$g_b = g_a + m_w$$

where m_w is the amount of feed water evaporated in kg per kg of dry air flowing through the spray chamber.

Applying the heat balance:

$$\begin{aligned} h_a + h_w &= h_b \\ &= (1.007t_b - 0.026) + g_b(2501 + 1.84t_b) \end{aligned} \quad (2.24)$$

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EXAMPLE 3.7

Moist air at a state of 21°C dry-bulb, 15°C wet-bulb (sling) and 101.325 kPa barometric pressure enters a spray chamber. If, for each kilogram of dry air passing through the chamber, 0.002 kg of water at 100°C is injected and totally evaporated, calculate the moisture content, enthalpy and dry-bulb temperature of the moist air leaving the chamber.

Answer

From CIBSE tables of psychrometric data,

$$h_a = 41.88 \text{ kJ per kg dry air}$$

$$g_a = 0.008 171 \text{ kg per kg dry air}$$

Since the feed water has a temperature of 100°C, its enthalpy is 419.06 kJ per kg of water injected, from CIBSE tables of properties of water at saturation.

Use of the equation for mass balance yields the moisture content of the moist air leaving the spray chamber:

$$\begin{aligned} g_b &= 0.008 171 + 0.002 \\ &= 0.010 171 \text{ kg per kg dry air} \end{aligned}$$

Use of the energy balance equation gives the enthalpy of the air leaving the chamber and hence, also, its dry-bulb temperature by equation (2.24)

$$\begin{aligned} h_b &= 41.88 + 0.002 \times 418.06 \\ &= 42.716 \text{ kJ per kg dry air} \\ &= (1.007t_b - 0.026) + 0.010 171(2501 + 1.84t_b) \end{aligned}$$

thus,

$$\begin{aligned} 42.716 &= 1.007t_b - 0.026 + 25.44 + 0.0187t_b \\ t_b &= 16.9^\circ\text{C} \end{aligned}$$

3.7 Steam injection

As in water injection, steam injection may be dealt with by a consideration of a mass and energy balance. If m_s kg of dry saturated steam are injected into a moving airstream of mass flow 1 kg of dry air per second, then we may write

$$g_b = g_a + m_s$$

and

$$h_b = h_a + h_s$$

If the initial state of the moist airstream and the condition of the steam is known, then the

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final state of the air may be determined, provided none of the steam is condensed. The change of state takes place almost along a line of constant dry-bulb temperature between limits defined by the smallest and largest enthalpies of the injected steam, provided the steam is in a dry, saturated condition. If the steam is superheated then, of course, the dry-bulb temperature of the airstream may increase by any amount, depending on the degree of superheat. The two limiting cases for dry saturated steam are easily considered. The lowest possible enthalpy is for dry saturated steam at 100°C. It is not possible to use steam at a lower temperature than this since the steam must be at a higher pressure than atmospheric if it is to issue from the nozzles. (Note, however, that steam could be generated at a lower temperature from a bath of warm water, which was not actually boiling.)

The other extreme is provided by the steam which has maximum enthalpy; the value of this is 2803 kJ kg⁻¹ of steam and it exists at a pressure of about 30 bar and a temperature of about 234°C.

What angular displacement occurs between these two limits, and how are they related to a line of constant dry-bulb temperature? These questions are answered by means of two numerical examples.

EXAMPLE 3.8

Dry saturated steam at 100°C is injected at a rate of 0.01 kg s⁻¹ into a moist airstream moving at a rate of 1 kg of dry air per second and initially at a state of 28°C dry-bulb, 11.9°C wet-bulb (sling) and 101.325 kPa barometric pressure. Calculate the leaving state of the moist airstream.

Answer

From psychrometric tables, $h_a = 33.11$ kJ per kg dry air,

$$g_a = 0.001\,937 \text{ kg per kg}$$

From NEL steam tables, $h_s = 2675.8$ kJ per kg steam.

$$\begin{aligned} g_b &= 0.001\,937 + 0.01 \\ &= 0.011\,937 \text{ kg per kg dry air, by the mass balance} \end{aligned}$$

$$\begin{aligned} h_b &= 33.11 + 0.01 \times 2675.8 \\ &= 59.87 \text{ kJ per kg dry air} \end{aligned}$$

Hence,

$$\begin{aligned} 59.87 &= (1.007t_b - 0.026) + 0.011937(2501 + 1.84t_b) \\ t_b &= 29.2^\circ\text{C} \end{aligned}$$

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EXAMPLE 3.9

Dry saturated steam with maximum enthalpy is injected at a rate of 0.01 kg s^{-1} into a moist airstream moving at a rate of 1 kg of dry air per second and initially at a state of 28°C dry-bulb, 11.9°C wet-bulb (sling) and 101.325 kPa barometric pressure. Calculate the leaving state of the moist airstream.

Answer

The psychrometric properties of state A are as for the last example. The moisture content at state B is also as in example 3.8.

From NEL steam tables, $h_s = 2803 \text{ kJ}$ per kg of steam, at 30 bar and 234°C saturated. Consequently,

$$\begin{aligned}h_b &= 33.11 + 0.01 \times 2803 \\ &= 61.14 \text{ kJ kg}^{-1} \\ g_b &= 0.001937 + 0.01 \\ &= 0.011937 \text{ kg kg}^{-1} \text{ dry air}\end{aligned}$$

Hence, as before:

$$\begin{aligned}61.14 &= (1.007t_b - 0.026) + 0.011937(2501 + 1.84t_b) \\ t_b &= 30.4^\circ\text{C}\end{aligned}$$

Figure 3.9 illustrates what occurs. It can be seen that for the range of states considered, the change in dry-bulb value is not very great. In fact, the angular displacement between the two condition lines for the last two examples of steam injection is only about 3 or 4 degrees. We can conclude that, although there is an increase in temperature, it is within the accuracy usually required in practical air conditioning to assume that the change of state following steam injection is up a line of constant dry-bulb temperature.

In the injection of superheated steam, every case should be treated on its merits, as the equation for the dry-bulb temperature resulting from a steam injection process shows:

$$t_b = \frac{h_b + 0.026 - 2501g}{1.007 + 1.84g} \quad (3.9)$$

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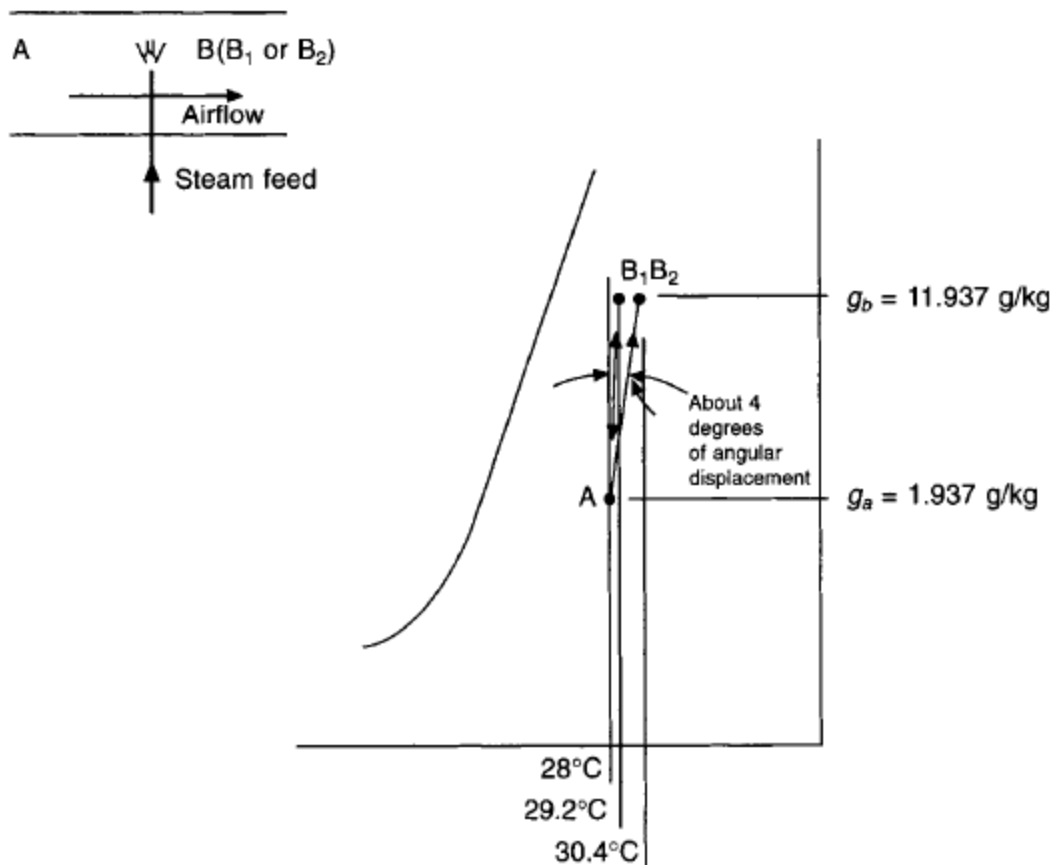


Fig. 3.9 A humidification process by the entire acceptance of injected dry saturated steam. See Examples 3.8 and 3.9.

3.8 Cooling and dehumidification with reheat

Figure 3.10(a) shows, in diagrammatic form, the sort of plant required. Moist air at a state A passes over the finned tubes of a cooler coil through which chilled water is flowing. The amount of dehumidification carried out is controlled by a dew point thermostat, C1, positioned after the coil. This thermostat regulates the amount of chilled water flowing through the coil by means of the three-way mixing valve R1. Air leaves the coil at state B, with a moisture content suitable for the proper removal of the latent heat gains occurring in the room being conditioned. The moisture content has been reduced from g_a to g_b and the cooler coil has a mean surface temperature of t_c , Figure 3.10(b) illustrating the psychrometric processes involved.

The cooling load is proportional to the difference of enthalpy between h_a and h_b , and the load on the heater battery is proportional to h_d minus h_b . It follows from this that part of the cooling load is being wasted by the reheat. This is unavoidable in the simple system illustrated, and is a consequence of the need to dehumidify first, and heat afterwards. It should be observed, however, that in general it is undesirable for such a situation to exist during maximum load conditions. Reheat is usually only permitted to waste cooling capacity under partial load conditions, that is, the design should be such that state B can adequately deal with both maximum sensible and maximum latent loads. These points are illustrated in the following example.

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EXAMPLE 3.10

Moist air at 28°C dry-bulb, 20.6°C wet-bulb (sling) and 101.325 kPa barometric pressure flows over a cooler coil and leaves it at a state of 10°C dry-bulb and 7.046 g per kg of dry air.

(a) If the air is required to offset a sensible heat gain of 2.35 kW and a latent heat gain of 0.31 kW in a space to be air-conditioned, calculate the mass of dry air which must be supplied to the room in order to maintain a dry-bulb temperature of 21°C therein.

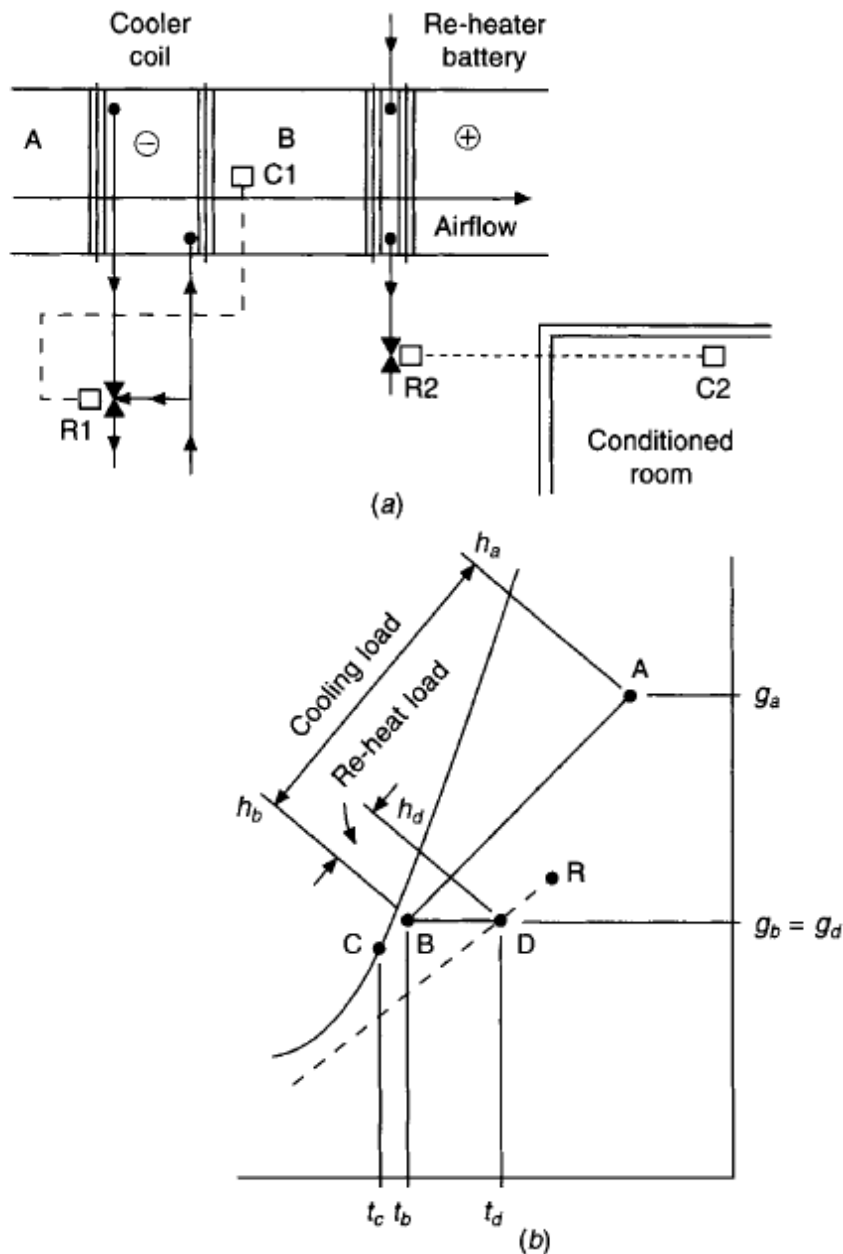


Fig. 3.10 (a) Plant arrangement for cooling and dehumidification with reheat. (b) Psychrometry for cooling and dehumidification with reheat.

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(b) What will be the relative humidity in the room?

(c) If the sensible heat gain diminishes by 1.175 kW but the latent heat gain remains unchanged, at what temperature and moisture content must the air be supplied to the room?

Answer

(a) If m_a kg per s of dry air are supplied at a temperature of 10°C , they must have a sensible cooling capacity equal to the sensible heat gain, if 21°C is to be maintained in the room.

Thus assuming the specific heat capacity of air is $1.012 \text{ kJ kg}^{-1} \text{ K}^{-1}$

$$m_a \times c_a \times (21^\circ - 10^\circ) = 2.35 \text{ kW}$$

then,

$$m_a = \frac{2.35}{1.012 \times 11^\circ} = 0.211 \text{ kg s}^{-1}$$

(b) 0.211 kg s^{-1} of dry air with an associated moisture content of 7.046 g per kg of dry air must take up the moisture evaporated by the liberation of 0.31 kW of latent heat. Assuming a latent heat of evaporation of, say, 2454 kJ per kg of water (at about 21°C), then the latent heat gain corresponds to the evaporation of

$$\frac{0.31 \text{ kJ s}^{-1}}{2454 \text{ kJ kg}^{-1}} = 0.0001263 \text{ kg of water per second}$$

The moisture associated with the delivery to the room of 0.211 kg s^{-1} of dry air will increase by this amount. The moisture picked up by each kg of dry air supplied to the room will be

$$\frac{0.1263}{0.211} = 0.599 \text{ g kg}^{-1}$$

Thus, the moisture content in the room will be equal to $7.046 + 0.599$, that is, 7.645 g kg^{-1} . The relative humidity at this moisture content and 21°C dry-bulb is found from tables of psychrometric data to be 49.3 per cent. If the relevant data are taken from a psychrometric chart, instead of from psychrometric tables, then similar but slightly different results are obtained, of adequate practical accuracy in most cases. The percentage saturation at 21°C dry-bulb is then just under 49 per cent.

(c) If 0.211 kg s^{-1} of dry air is required to absorb only 1.175 kW of sensible heat then, if 21°C is still to be maintained, the air must be supplied at a higher temperature:

$$\text{temperature rise} = \frac{1.175}{1.012 \times 0.211} = 5.5 \text{ K}$$

Thus, the temperature of the supply air is 15.5°C .

Since the latent heat gains are unrelated, the air supplied to the room must have the same ability to offset these gains as before; that is, the moisture content of the air supplied must still be 7.046 g per kg of dry air.

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EXAMPLE 3.11

(a) For sensible and latent heat gains of 2.35 and 0.31 kW, respectively, calculate the load on the cooler coil in example 3.10.

(b) Calculate the cooling load and the reheater load for the case of 1.175 kW sensible heat gains and unchanged latent gains.

Answer

(a) The load on the cooler coil equals the product of the mass flow of moist air over the coil and the enthalpy drop suffered by the air. Thus, as an equation:

$$\text{cooling load} = m_a \times (h_a - h_b)$$

The notation adopted is the same as that used in Figure 3.10. In terms of units, the equation can be written as

$$\text{kW} = \frac{\text{kg of dry air}}{\text{s}} \times \frac{\text{kJ}}{\text{kg of dry air}}$$

Using enthalpy values obtained from tables for the states A and B, the equation becomes

$$\begin{aligned} \text{cooling load} &= 0.211 \times (59.06 - 27.81) \\ &= 0.211 \times 31.25 \\ &= 6.59 \text{ kW} \end{aligned}$$

(b) Since it is stipulated that the latent heat gains are unchanged, air must be supplied to the conditioned room at the same state of moisture content as before. That is, the cooler coil must exercise its full dehumidifying function, the state of the air leaving the coil being the same as in example 3.11(a), above. Hence the cooling load is still 6.59 kW. To deal with the diminished sensible heat gains it is necessary to supply the air to the room at a temperature of 15.5°C, as was seen in example 3.10(c). The air must, therefore, have its temperature raised by a reheater battery from 10°C to 15.5°C. Since the moisture content is 7.046 g kg⁻¹ at both these temperatures, we can find the corresponding enthalpies at states B and D directly from a psychrometric chart or by interpolation (in the case of state D) from psychrometric tables.

Hence, we can form an equation for the load on the reheater battery:

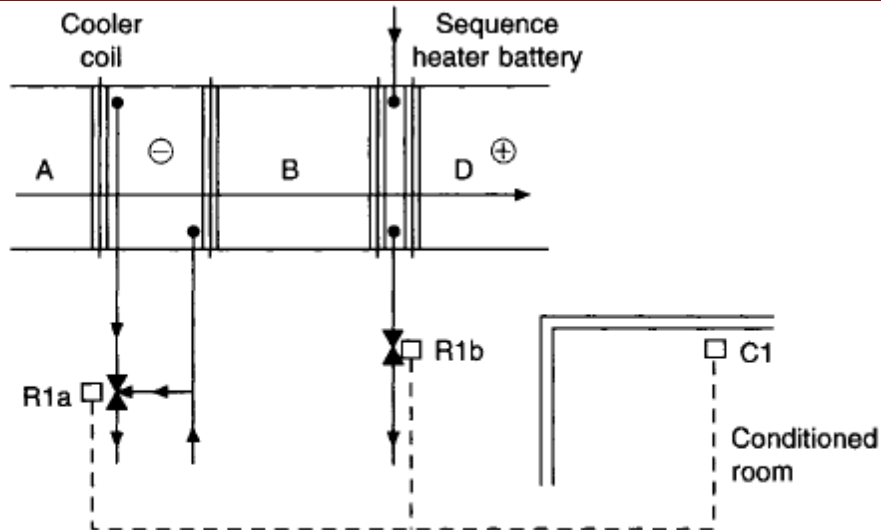
$$\text{reheater load} = 0.211 \times (33.41 - 27.81) = 1.18 \text{ kW}$$

EXAMPLE 3.12

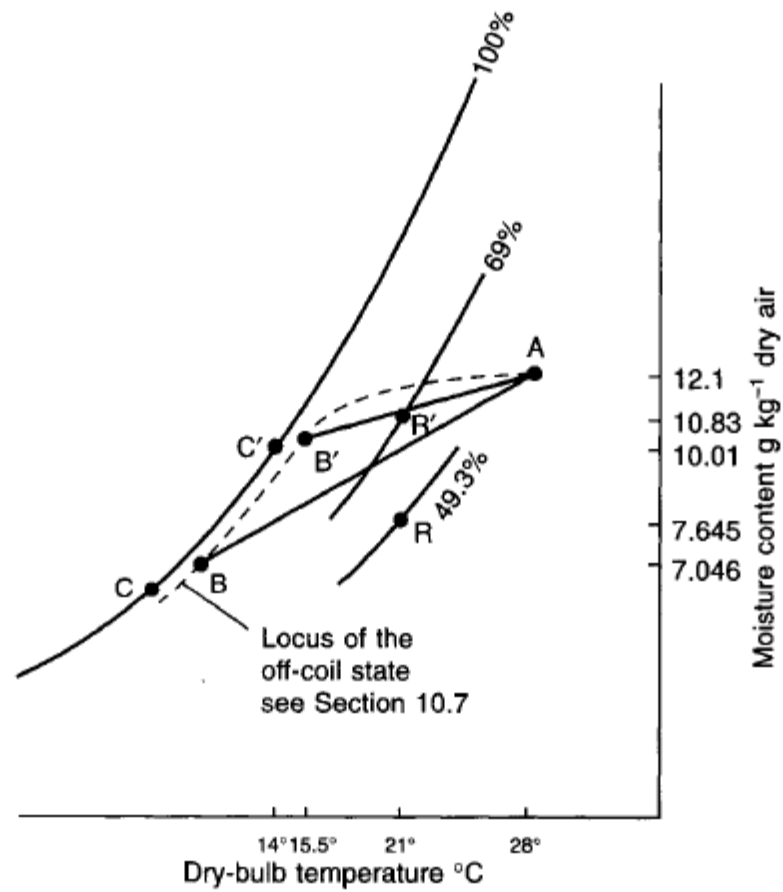
If the plant in example 3.10 is arranged for sequence control over the cooler coil and heater battery, and if under the design conditions mentioned in that example the plant is able to maintain 21°C dry-bulb and 49.3 per cent relative humidity in the conditioned space, calculate the relative humidity that will be maintained there, under sequence control, if the sensible heat gains diminish to 1.175 kW, the latent gains remaining unaltered.

It is given that the mean coil surface temperature under the condition of partial load is 14°C. (See section 10.7 on the partial load performance of cooler coils.)

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(a)



(b)

Fig. 3.11 (a) Plant arrangement for heating and cooling in sequence. (b) Psychrometry for heating and cooling in sequence, for Example 3.12.

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Answer

As was mentioned in section 3.4, the mean coil surface temperature (identified by the state point C), lies on the straight line joining the cooler coil 'on' and 'off' states (points A and B) where it cuts the saturation curve. Figure 3.11(b) illustrates this, and also shows that the moisture content of state B' is greater than that of B. It is possible to calculate the moisture content of the supply air by assuming that the dry-bulb scale on the psychrometric chart is linear. By proportion, we can then assess a reasonably accurate value for the moisture content of the air leaving the cooler coil under the partial load condition stipulated in this example:

$$\frac{g_a - g_{b'}}{g_a - g_{c'}} = \frac{t_a - t_{b'}}{t_a - t_{c'}}$$

From tables, or from a psychrometric chart, $g_a = 12.10 \text{ g kg}^{-1}$ and $g_{c'} = 10.01 \text{ g kg}^{-1}$. Also, it was calculated in example 3.10(b) that the supply temperature must be 15.5°C dry-bulb for the partial sensible load condition.

Hence we can write

$$\begin{aligned} g_{b'} &= g_a - (g_a - g_{c'}) \times \frac{(t_a - t_{b'})}{(t_a - t_{c'})} \\ &= 12.10 - (12.10 - 10.01) \times \frac{(28 - 15.5)}{(28 - 14)} = 10.23 \text{ g kg}^{-1} \end{aligned}$$

Since the moisture pick-up has been previously calculated as 0.599 g kg^{-1} , the moisture content in the conditioned room will be 10.83 g kg^{-1} ; at 21°C dry-bulb the relative humidity will be about 69 per cent.

3.9 Pre-heat and humidification with reheat

Air conditioning plants which handle fresh air only may be faced in winter with the task of increasing both the moisture content and the temperature of the air they supply to the conditioned space. Humidification is needed because the outside air in winter has a low moisture content, and if this air were to be introduced directly to the room there would be a correspondingly low moisture content there as well. The low moisture content may not be intrinsically objectionable, but when the air is heated to a higher temperature its relative humidity may become very low. For example, outside air in winter might be at -1°C , saturated (see chapter 5). The moisture content at this state is only 3.484 g per kg of dry air. If this is heated to 20°C dry-bulb, and if there is a moisture pick-up in the room of 0.6 g per kg of dry air, due to latent heat gains, then the relative humidity in the room will be as low as 28 per cent. This value may sometimes be seen as too low for comfort. The plant must increase the temperature of the air, either to the value of the room temperature if there is background heating to offset fabric losses, or to a value in excess of this if it is intended that the air delivered should deal with fabric losses.

The processes whereby the moisture content of air may be increased were discussed in section 3.5. The actual method chosen depends on the application but, for comfort air

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conditioning or for any application where people are present in the conditioned space, the use of an air washer or any method involving the recirculation of spray water, or the exposure of a wetted surface area to the airstream, is not recommended. This is because of the risk to health caused by the presence of micro-organisms in the water. The use of dry steam injection is much preferred.

As an exercise in psychrometry, it is worth considering the case where an air washer is used. If 100 per cent fresh air is handled in cold weather it is pre-heated, passed through an air washer where it undergoes adiabatic saturation, and reheated to the temperature at which it must be supplied to the room. Pre-heating and adiabatic saturation permit the relative humidity in the room to be controlled, and reheating allows the temperature therein to be properly regulated, in winter.

Figure 3.12(a) shows, in a diagrammatic form, a typical plant. Opening the modulating valve R1 in the return pipeline from the pre-heater increases the heating output of the battery and provides the necessary extra energy for the evaporation of more water in the washer, if an increase in the moisture content of the supply air is required. Similarly, opening the control valve R2, associated with the reheater, allows air at a higher temperature to be delivered to the room being conditioned. C1 and C2 are a room humidistat and a room thermostat, respectively.

EXAMPLE 3.13

The plant shown in Figure 3.12(a) operates as illustrated by the psychrometric changes in Figure 3.12(b).

Air is pre-heated from -5.0°C dry-bulb and 86 per cent saturation to 23°C dry-bulb. It is then passed through an air washer having a humidifying efficiency of 85 per cent and using recirculated spray water. Calculate the following:

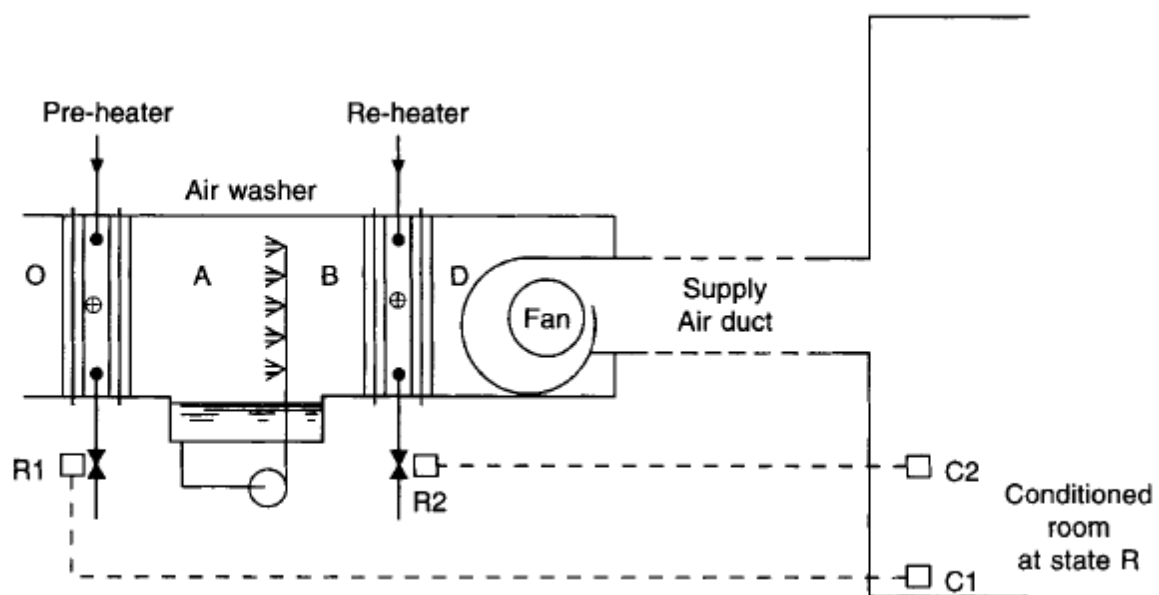
- (a) The relative humidity of the air leaving the washer
- (b) The cold water make-up to the washer in litres s^{-1} , given that the airflow rate leaving the washer is $2.5 \text{ m}^3 \text{ s}^{-1}$
- (c) The duty of the pre-heater battery in kW
- (d) The temperature of the air supplied to the conditioned space if the sensible heat losses from it are 24 kW and 20°C dry-bulb is maintained there
- (e) The duty of the reheater battery
- (f) The relative humidity maintained in the room if the latent heat gains therein are 5 kW

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Answer

(a) At state O the moisture content is found from tables to be 2.137 g kg^{-1} . Consequently, at 23°C dry-bulb and 2.137 g kg^{-1} the wet-bulb (sling) value is found from tables to be 10°C . This state is represented by the point A in the diagram. Assuming that the process of adiabatic saturation occurs up a wet-bulb line, we can establish from tables that the moisture content at state C, the apparatus dew point, is 7.659 g kg^{-1} . If the further assumption is made that the dry-bulb scale is linear, we can evaluate the dry-bulb temperature at state B, leaving the washer, by proportion:

$$\begin{aligned}t_b &= 23 - 0.85(23 - 10) \\ &= 23 - 11.05 \\ &= 12^\circ\text{C dry-bulb}\end{aligned}$$



(a)

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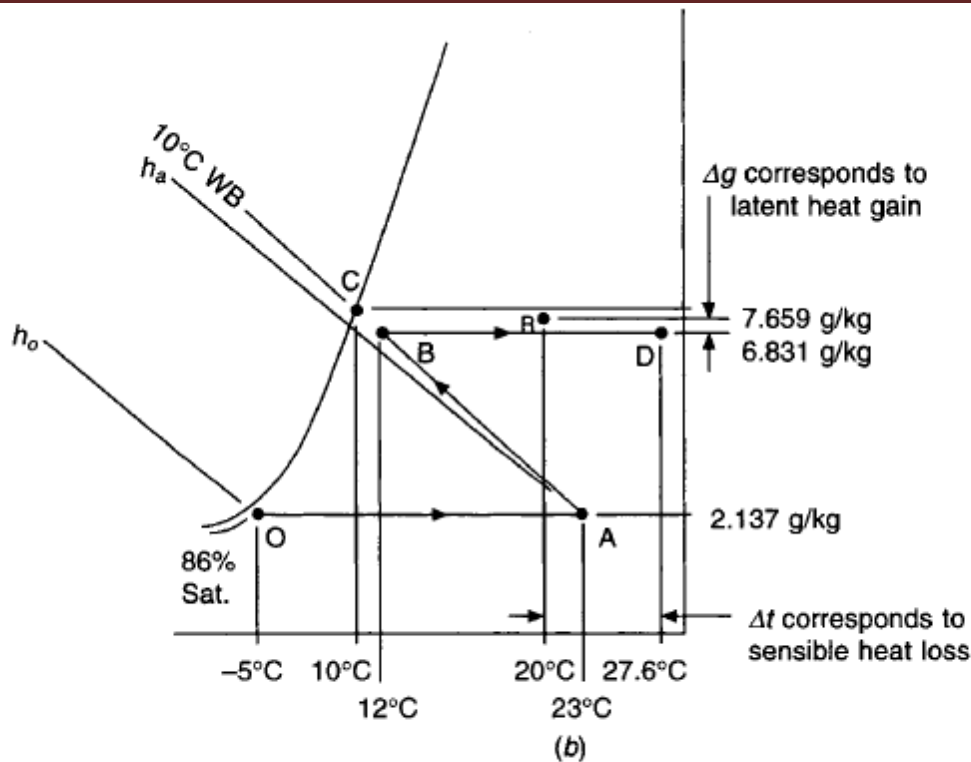


Fig. 3.12 (a) Plant arrangement for pre-heating 100% fresh air with adiabatic saturation and reheat.
(b) The psychrometry for Example 3.13.

We know that the dry-bulb temperature at the point C is 10°C because the air is saturated at C and it has a wet-bulb temperature of 10°C.

The above answer for the dry-bulb at B is approximate (because the dry-bulb scale is not linear) but accurate enough for all practical purposes. On the other hand, we may calculate the moisture content at B with more exactness, since the definition of humidifying efficiency is in terms of moisture content changes and this scale is linear on the psychrometric chart.

$$\begin{aligned} g_b &= 7.659 - (1 - 0.85) \times (7.659 - 2.137) \\ &= 7.659 - 0.828 \\ &= 6.831 \text{ g kg}^{-1} \end{aligned}$$

We can now refer to tables or to a chart and determine that the relative humidity at 12°C dry-bulb and 6.831 g kg⁻¹ is 78 per cent. Percentage saturation is virtually the same as relative humidity. More accurately, we might perhaps determine the humidity at a state of 10°C wet-bulb (sling) and 6.831 g kg⁻¹. The yield in accuracy is of doubtful value and the method is rather tedious when using tables.

(b) The cold water make-up depends on the mass flow of dry air. It may be determined that the humid volume of the air leaving the washer at state B is 0.8162 m³ kg⁻¹ of dry air. The cold water fed from the mains to the washer serves to make good the losses due to evaporation within the washer. Since the evaporation rate is (6.831 - 2.137) g kg⁻¹, we may calculate the make-up rate:

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$$\begin{aligned}\text{make-up rate} &= \frac{(6.831 - 2.137) \times 2.5}{1000 \times 0.8162} \\ &= 0.0144 \text{ kg s}^{-1} \text{ or litres s}^{-1}\end{aligned}$$

(c) The pre-heater must increase the enthalpy of the air passing over it from h_o to h_a . Reference to tables establishes that $h_o = 0.298 \text{ kJ kg}^{-1}$ and $h_a = 28.57 \text{ kJ kg}^{-1}$. Then,

$$\begin{aligned}\text{pre-heater duty} &= \frac{2.5 \times (28.57 - 0.298)}{0.8162} \\ &= 3.063 \text{ kg s}^{-1} \times 28.27 \text{ kJ s}^{-1} \\ &= 86.59 \text{ kW}\end{aligned}$$

(d) 3.063 kg s^{-1} of air diffuses throughout the room and its temperature falls from t_d to 20°C as it offsets the heat loss. Assuming the specific heat capacity of water vapour is $1.89 \text{ kJ kg}^{-1} \text{ K}^{-1}$ and that of dry air $1.012 \text{ kJ kg}^{-1} \text{ K}^{-1}$, a heat balance equation can be written:

$$\begin{aligned}\text{Sensible heat loss} &= (\text{mass flow rate of dry air} \times \\ &\quad \text{specific heat of dry air} + \text{associated} \\ &\quad \text{moisture} \times \text{specific heat of water vapour}) \\ &\quad \times (\text{supply air temperature} - \text{room temperature}) \\ 24 &= 3.063 \times (1.012 + 0.006831 \times 1.89) \times (t_d - 20) \\ &= 3.063 \times 1.025 \times (t_d - 20)\end{aligned}$$

whence $t_d = 27.6^\circ\text{C}$

The term involving the specific heat of dry air plus the moisture content times the specific heat of water vapour, is termed the specific heat of humid air or the humid specific heat. In this case its value is $1.025 \text{ kJ kg}^{-1} \text{ K}^{-1}$.

(e) The reheater battery thus has to heat the humid air from the state at which it leaves the washer to the state at which it enters the room, namely from 12°C dry-bulb and 6.831 g kg^{-1} to 27.6° dry-bulb and 6.831 g kg^{-1} .

$$\begin{aligned}\text{Reheater duty} &= 3.063 \times 1.025 \times (27.6 - 12) \\ &= 49 \text{ kW}\end{aligned}$$

Alternatively, we can determine the enthalpies of the states on and off the heater, by interpolating from tables or reading directly from a psychrometric chart:

$$\begin{aligned}\text{reheater duty} &= 3.063 \times (45.2 - 29.3) \\ &= 48.7 \text{ kW}\end{aligned}$$

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(f) A mass balance must be struck to determine the rise in moisture content in the room as a consequence of the evaporation corresponding to the liberation of the latent heat gains:

$$\begin{aligned} \text{latent heat gain in kW} &= (\text{kg of dry air per hour delivered to the room}) \\ &\times (\text{the moisture pick-up in kg of water per kg of dry air}) \\ &\times (\text{the latent heat of evaporation of water in} \\ &\hspace{15em} \text{kJ per kg of water}) \end{aligned}$$

$$5.0 = 3.063 \times (g_r - 0.006\ 831) \times 2454$$

whence

$$g_r = 0.007\ 496 \text{ kg per kg dry air}$$

From tables or from a chart it may be found that at a state of 20°C dry-bulb and 7.496 g per kg dry air, the relative humidity is about 51 per cent and, for use in Example 3.14(b), $h_r = 39.14 \text{ kJ kg}^{-1}$.

3.10 Mixing and adiabatic saturation with reheat

Figure 3.13(a) shows a plant arrangement including an air washer which, although undesirable for comfort air conditioning, is retained here to illustrate the psychrometry involved. Air at state R is extracted from a conditioned room and partly recirculated, the remainder being discharged to atmosphere. The portion of the extracted air returned to the air conditioning plant mixes with air at state O, drawn from outside, and forms a mixture state M. The air then passes through an air washer, the spray water of which is only recirculated and adiabatic saturation occurs, the state of the air changing from M to W (see Figure 3.13(b)) along a line of constant wet-bulb temperature (see sections 3.5 and 3.6). An extension of the line MW cuts the saturation curve at a point A, the apparatus dew point. To deal with a particular latent heat gain in the conditioned room it is necessary to supply the air to the room at a moisture content g_s , it being arranged that the difference of moisture content $g_r - g_s$, in conjunction with the mass of air delivered to the room, will offset the latent gain. In other words, the air supplied must be dry enough to absorb the moisture liberated in the room.

It is evident that the moisture content of the air leaving the washer must have a value g_w , equal to the required value, g_s . This is amenable to calculation by making use of the definition of the effectiveness of an air washer, in terms of g_a , g_w and g_m (see section 3.5).

EXAMPLE 3.14

If the room mentioned in example 3.13 is conditioned by means of a plant using a mixture of recirculated and fresh air, of the type illustrated in Figure 3.13(a), calculate:

- the percentage of the air supplied to the room by mass which is recirculated, and
- the humidifying efficiency of the air washer.

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Answer

(a) Since the wet-bulb scale is not linear, it is not accurate enough to calculate the mixing proportions on this basis. Instead, one must make use of changes of enthalpy or moisture content. Bearing in mind that lines of constant enthalpy are not parallel to lines of constant wet-bulb temperature, some slight inaccuracy is still present if the assumption is made that the change of state accompanying a process of adiabatic saturation is along a line of constant enthalpy. However, this is unavoidable, and so such an assumption is made.

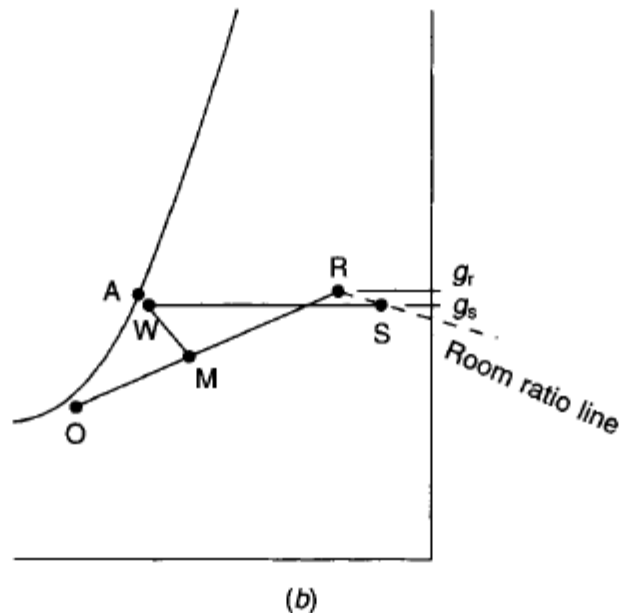
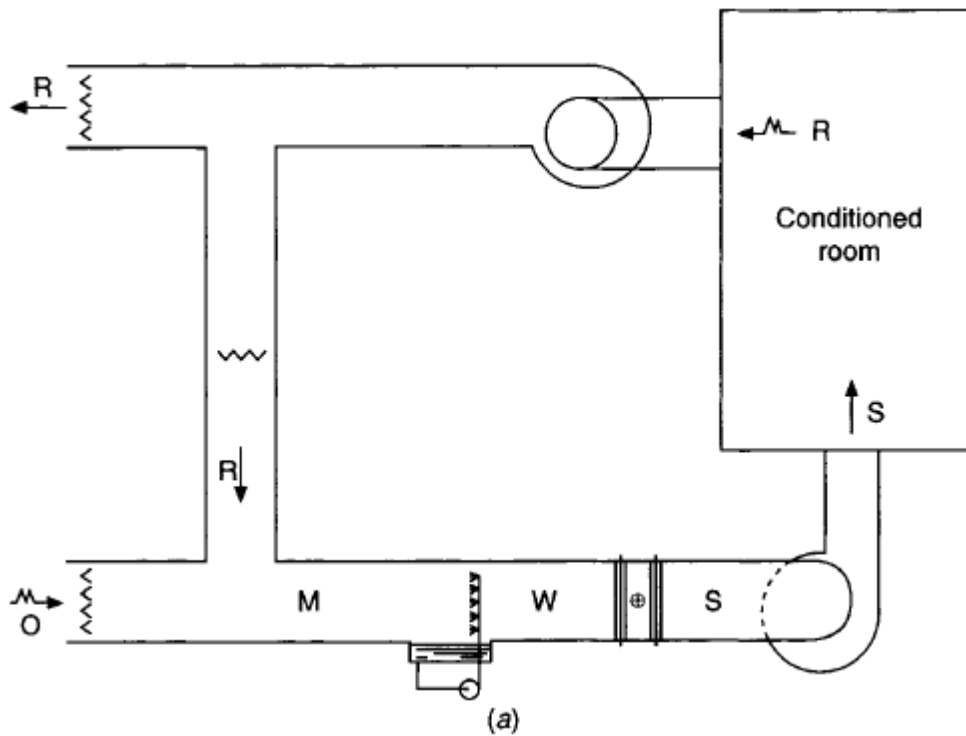


Fig. 3.13 (a) Plant arrangement to permit variable mixing proportions of fresh and recirculated air.
(b) Psychrometry for example 3.14.

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Referring to Figure 3.13(b) it can be seen that

$$h_a \approx h_w \approx h_m$$

From tables and Figure 3.12(b) it is established that h_w (at 12°C dry-bulb and 6.831 g kg⁻¹) is 29.30 kJ kg⁻¹.

From the principles set out in section 3.2, governing the change of state associated with a mixing process, it is clear that the percentage of recirculated air, by mass,

$$\begin{aligned} &= \frac{h_m - h_o}{h_r - h_o} \times 100 \\ &= \frac{29.30 - 0.298}{39.14 - 0.298} \times 100 \\ &= 75 \text{ per cent} \end{aligned}$$

Thus, 75 per cent of the air supplied to the room, if recirculated and mixed with 25 per cent of air from outside, will have an enthalpy of 29.30 kJ kg⁻¹ and a wet-bulb of 10°C (sling). If adiabatic saturation is then to produce a state of 12°C dry-bulb and 6.831 g kg⁻¹, the humidifying efficiency of the washer used can no longer be the value used for example 3.13, namely, 85 per cent. An entirely different washer must be used for the above calculations to be valid and this must have an effectiveness which may be calculated as follows:

(b) Since efficiency is expressed in terms of moisture content, it is necessary to determine the value of g_m , the values of g_a and g_w being already known.

$$\begin{aligned} g_m &= 0.75g_r + 0.25g_o \\ &= 0.75 \times 7.497 + 0.25 \times 2.137 \\ &= 6.157 \text{ g kg}^{-1} \end{aligned}$$

$$\begin{aligned} \text{Humidifying efficiency} &= \frac{g_w - g_m}{g_a - g_m} \times 100 \\ &= \frac{6.831 - 6.157}{7.659 - 6.157} \times 100 \\ &= 45 \text{ per cent} \end{aligned}$$

In practical terms, this is a low efficiency.

If the washer used in this example had an efficiency of 85 per cent, as in example 3.13, then the calculations would not have been so easy. The line AWM would have had to have been at a lower wet-bulb value in order to fulfil two requirements:

- (i) $g_w = 6.831 \text{ g kg}^{-1}$
- (ii) $\frac{g_w - g_m}{g_a - g_m} \times 100 = 85 \text{ per cent}$

For this to be the case, the dry-bulb temperature of W must obviously be less than 10°C. The easiest way to achieve a practical solution is by drawing a succession of lines representing processes of adiabatic saturation on a psychrometric chart and calculating several values of efficiency until one of acceptable accuracy is obtained.

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3.11 The use of dry steam for humidification

It is only if the air can accept the additional moisture, that humidification may be achieved by the injection of spray water or dry steam. If the state of the air into which it is proposed to inject moisture is saturated, or close to saturation, some or all of the moisture added will not be accepted and will be deposited downstream in the air handling plant or the duct system. It is therefore essential to ensure that the airstream is sufficiently heated prior to moisture injection.

EXAMPLE 3.15

A room is to be maintained at a state of 20°C dry-bulb and 50 per cent saturation by a plant handling 0.5 m³ s⁻¹ of outside air at a state of -2°C saturated. The airstream is heated to a temperature warm enough to offset a heat loss of 2.5 kW and dry steam is then injected to maintain the humidity required in the room. Calculate the supply air temperature and the heating and humidification loads. See Figure 3.14.

Answer

From psychrometric tables or a psychrometric chart, air at the outside state has a moisture content of 3.205 g kg⁻¹, an enthalpy of 5.992 kJ kg⁻¹ and a specific volume of 0.7716 m³ kg⁻¹. Assuming specific heats of 1.012 and 1.89 kJ kg⁻¹ K⁻¹, respectively, for dry air and water vapour, the humid specific heat of the fresh air handled is

$$1 \times 1.012 + 0.003\ 205 \times 1.89 = 1.018 \text{ kJ kg}^{-1} \text{ K}^{-1}$$

The heat loss is offset by the supply of air at a temperature t_s , warmer than the room temperature of 20°C. Hence

$$2.5 = (0.5/0.7716) \times 1.018 \times (t_s - 20)$$

whence $t_s = 23.8^\circ\text{C}$.

From tables or a chart the enthalpy at state B, leaving the heater battery, is 32.10 kJ kg⁻¹. See Figure 3.14. Hence the heater battery load is

$$(0.5/0.7716) \times (32.10 - 5.992) = 16.92 \text{ kW}$$

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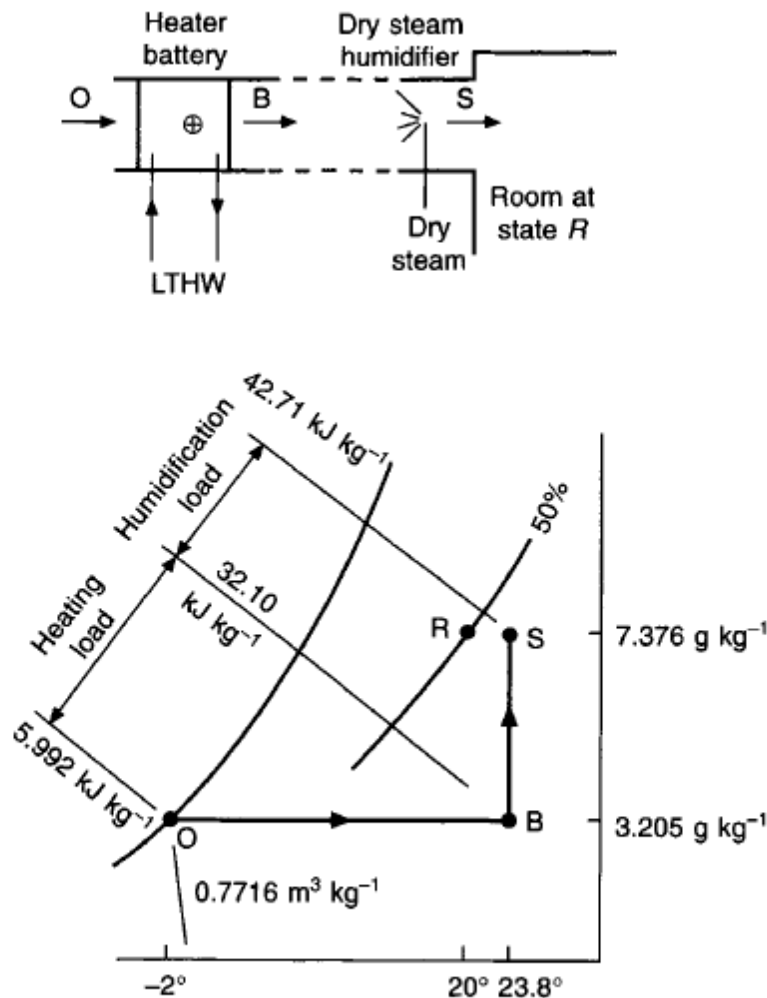


Fig. 3.14 The use of dry steam for humidification. See Example 3.15.

Alternatively

$$(0.5/0.7716) \times 1.018 \times (23.8 + 2) = 17.02 \text{ kW}$$

The first method is preferred because enthalpy values are based on well-established thermodynamic properties of dry air and water vapour whereas, on the other hand, the values of the specific heats used for the second method may not be exactly correct.

Assuming the change of state resulting from the injection of dry saturated steam is up a dry-bulb line, the state of the air supplied to the room is 23.8°C dry-bulb and 7.376 g kg⁻¹, at which the enthalpy is determined as 42.71 kJ kg⁻¹ from tables or a chart. Hence the humidification load is

$$(0.5/0.7716) \times (42.71 - 32.10) = 6.88 \text{ kW}$$

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Exercises

1. Dry saturated steam at 1.2 bar and 104.8°C is injected at a rate of 0.01 kg s⁻¹ into an airstream having a mass flow rate of dry air of 1 kg s⁻¹ and an initial state of 28°C dry-bulb, 11.9°C wet-bulb (sling). Calculate the leaving state of the airstream if all the injected steam is accepted by it. Use equation (2.24) as necessary.

Answers

29.3°C, 11.937 g kg⁻¹.

2. Air at 40°C dry bulb temperature and 30% relative humidity is passed through an adiabatic air washer at the rate of 28 m³/min. If the effectiveness of the air washer is 80%, find the conditions of the air leaving the air washer, and the amount of water vapor added to the air per minute.

Answer

$W_2 = 0.019 \text{ kg/kg}_{d.a.}$, $t_{d2} = 28^\circ\text{C}$, $\phi_2 = 80\%$, $m_s = 0.154 \text{ kg/min}$

3. An air conditioning system controls the indoor conditions of a room at 20°C dry bulb temperature and 50% relative humidity, when the outdoor conditions are at 5°C dry bulb temperature and 70% relative humidity. 2 m³/s of recycled air from the room is mixing with 1 m³/s of the fresh air, the mixture then flows across a heater and come out of it at 27°C, and supply to the room. Sketch the process on the psychrometric chart and determine:

(a) the amount of heat that gives the heater to the air in kW.

(b) the heating load of the room in kW.

Answer

45.13 kW, 14.56 kW

4. Air at 28°C dry bulb temperature and 50% relative humidity passes through a cooling coil at a rate of 1 m³/s, the air leaves the coil at 10°C dry bulb temperature and 0.007 kg/kg_{d.a.} moisture content. Calculate the following by using the psychrometric chart.

(a) mass flow rate of the air (b) total heat removed from the air

(c) contact factor of the coil (d) rate of condensing of water vapor

Answer: 1.150 kg/s, 34.5 kW, 0.81, 20.286 kg/h.

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5. $0.95 \text{ m}^3/\text{s}$ of outside air passes over a cooling coil. The outdoor air is at 35°C dry bulb and 25°C wet bulb temperature, and the conditioned space is being maintained at 26°C dry bulb temperature and 45% relative humidity. the sensible heat ratio is calculated as 0.72. the air leaving the coil is 90% saturated.

(a) find the apparatus dew point, and the temperature of the air leaving the coil

(b) how much cooling in kW is the unit doing

(c) how much moisture in $\text{kg}/\text{kg}_{\text{d.a}}$ is condensed out of the incoming air per hour.

Answer

(a) 8°C , 10.4°C (b) 51.39 kW (c) 34.47 kg/h

6. The sensible heat gain of a room is 4.8 kW and its latent heat gain is 1.4 kW. A conditioned air supply of $0.5 \text{ m}^3/\text{s}$ is to be delivered to the room. If the room is to be maintained at 25°C DBT, find the relative humidity that will result in the conditioned room if the air supply is 17°C and 90% RH.

Answer

60%

7. Air enters a chamber at 10°C DBT and 5°C WBT at a rate of $100 \text{ m}^3/\text{min}$. The barometer reads a pressure of 1.01325 bar. While passing through the chamber, the air absorbs sensible heat at the rate of 40 kW and picks up 45 kg/h of saturated steam at 105°C . Determine the dry and wet bulb temperatures of the air leaving the chamber.

Answer

26.5°C , 18.1°C

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Thermal Comfort

One of the main purposes of an air-conditioning system is to provide conditions for human thermal comfort. Comfort is related to levels of optimum acceptability which has been established by response tests for all types of subjects under varying environmental conditions of temperature and humidity. It is a subjective quality, personal to individuals depending on sex, age, state of health, clothing, environmental conditions, and very often depending on individual preferences. In general, comfort occurs when body temperatures are held within narrow ranges, and skin moisture is low. The level of noise in the controlled environment also affects the feeling of comfort.

4.1 Heat Balance Equation

The physical basis of comfort lies in the thermal balance of the body, i.e. the heat produced by the body's metabolism must be dissipated to the environment, otherwise the body would overheat.

The total energy production rate of the body is the sum of the production rates of heat Q and work W and can be written in the form

$$Q + W = MA_{\text{skin}} \dots\dots\dots (4.1)$$

Where

M is the rate of metabolic energy production per unit surface area

A_{skin} is the total surface area of skin.

The thermal balance of the body can be expressed by the equation,

$$S = (M - W) - E \pm R \pm C \dots\dots\dots (4.2)$$

Where

$(M - W)$ is the net surplus heat to be liberated or stored (metabolic rate minus the useful rate of working)

E is the heat loss by evaporation

R is the heat gain or loss by radiation

C is the heat gain or loss by convection

S is the rate at which heat is stored within the body.

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Under steady state conditions, the body remains comfortable and healthy because S is zero. In an oppressively hot environment, the load imposed upon E , R and C may be so great that S is positive and the body temperature will rise, eventually resulting in heat stroke.

Heat dissipation from the body (Table 4.1) to immediate surroundings occurs by several modes of heat exchange:

- Sensible heat flow from the skin
- Latent heat flow from evaporation of sweat and from evaporation of moisture diffused through the skin
- Sensible heat flow during respiration
- Latent heat flow due to evaporation of moisture during respiration.

Table 4.1 Heat output of the body in various activities

Activity	Watts
Sleeping	min. 70
Sitting, moderate movement, e.g. typing on computer	160 – 190
Sitting, heavy arm and leg movements	190 – 230
Standing, moderate work, some walking	220 – 290
Walking, moderate lifting or pushing	290 – 410
Intermittent heavy lifting, digging	440 – 580
Hard, sustained work	580 – 700

Sensible and latent heat losses from the skin are typically expressed in terms of environmental factors, skin temperature, and skin wettedness. The main independent environmental variables can be summarized as air temperature, mean radiant temperature and relative air velocity and ambient water vapor pressure.

4.2 Thermal Interchange with Environment

The human body is continually gaining and producing heat as well as losing heat to its surroundings to maintain temperature equilibrium. Body heat gains come from two source:

- Heat produced within the body itself as a result of metabolic processes.

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- Heat gained by body from external sources, by radiation from the sun or other hot objects or surfaces, and by convection from the surrounding air.

Heat is lost from the body by:

(a) **Conduction:** Heat loss by conduction depends on the temperature difference between the body surface and the object with which the body is in direct contact. Heat lost by conduction from the body can be neglected as the amount of body surface in contact with an external surface is usually too small and the period of contact is short too.

(b) **Convection** (about 30%): Heat loss due to convection takes place from the body to the air in contact with the skin or clothing. The rate of convection heat loss is increased by a faster rate of air movement, by a lower air temperature and a higher skin temperature.

(c) **Radiation** (about 45%): Radiant heat loss depends on the temperature of the body surface and the temperature of the opposing surfaces. Thus the human body will radiate heat to walls, ceilings, floors, windows, and to the out of doors if these surfaces are at a lower temperature than the body surface. Conversely, the body gains by radiation from the sun or from any surface warmer than the skin surface. Body skin temperature ranges between 30°C and 34°C with an average of 32.2°C for a healthy person engaged in light activity.

(d) **Evaporation** (about 25%): Heat loss by evaporation is governed by the rate of evaporation, which in turn depends on the humidity of air (the dryer the air, the faster the evaporation) and on the amount of moisture available for evaporation.

Metabolic heat generation

In choosing optimal conditions for comfort and health, knowledge of the energy expended during the course of routine physical activities is necessary, since heat production increases in proportion to exercise intensity. The unit used to measure the metabolic rate is *met*. One met represents the average heat produced by a sedentary average person at normal mean radiant temperature, i.e. 1 met = 58.2 W/m². Table 4.2 lists the typical metabolic heat generation for various activities.

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Table 4.2 Typical metabolic heat generation for various activities

Activities	W/m ²	met
Resting		
Sleeping	40	0.7
Reclining	45	0.8
Seated, quiet	60	1.0
Standing, relaxed	70	1.2
Walking		
3.2 km/h (0.9 m/s)	115	2.0
4.3 km/h (1.2 m/s)	150	2.6
6.4 km/h (1.8 m/s)	220	3.8
Office activities		
Reading, seated	55	1.0
Writing	60	1.0
Typing	65	1.1
Filing, seated	70	1.2
Filing, standing	80	1.4
Walking about	100	1.7
Lifting/packing	120	2.1

Clothing affects comfort, since it acts as an insulation. The unit measuring the insulating effect of clothing on a human subject is *clo*, where, $1 \text{ clo} = 0.155 \text{ km}^2/\text{W}$.

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4.3 Environmental Parameters and Indices

Environmental parameters

Environmental parameters that affect human comfort can be categorized into (a) directly measured parameters and (b) calculated parameters.

The following are the frequently used directly measured psychrometric parameters:

- Dry bulb temperature
- Wet bulb temperature
- Dew point temperature
- Water vapor pressure
- Total atmospheric pressure
- Relative humidity
- Humidity ratio
- Air velocity

The mean radiant temperature is derivable and, hence, a *calculated parameter*. It is the temperature of a uniform black enclosure in which a solid body or occupant would exchange the same amount of radiant heat as in the existing non-uniform environment. Fanger identified two additional calculated parameters, which are *activity level* and *clothing*. In addition to the above, the other secondary factors such as day-to-day temperature variation, age, adaptability, sex, etc. also influence comfort.

Environmental indices

An environmental index combines two or more parameters, such as air temperature, mean radiant temperature, humidity or air velocity into a single variable. The *effective temperature* (ET^*) is probably the most common environmental index and has the widest range of applications.

The *effective temperature* (ET^*) is defined as the dry bulb temperature of a uniform enclosure at 50% RH in which humans would have the same net heat exchange by radiation, convection, and evaporation as they would in the varying humidities of the test environment.

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Another approach used to evaluate the combined effect of temperature and humidity is the *Heat Stress Index*. This index is the ratio of the total evaporative heat loss required for thermal equilibrium to the maximum evaporative heat loss possible for the environment, multiplied by 100 for steady-state conditions (skin temperature is held constant at 35°C in order to limit the rise in body temperature, the sweat rate should not exceed one liter per hour to limit the loss of body fluid). The heat stress index is therefore defined as

$$\text{Heat Stress Index} = Q_E / Q_{E,\max} \dots\dots\dots (4.3)$$

Where

Q_E is the actual evaporative loss

$Q_{E,\max}$ is the maximum evaporative heat loss with the skin temperature at 35°C.

4.4 Comfort Charts

In identical environments, different people perceive comfort in different ways. In the same built environment, some may feel chilly while others may feel warm. Dry bulb temperature is not a reliable indication of how warm or cold an occupant will feel in a room. The effects of both relative humidity and air velocity need also to be considered.

In the same context, ASHRAE and other researchers have conducted extensive research over the years to relate the above factors to human comfort. From the results of these tests emerged the concept of an *effective temperature*. This index is a measure of comfort which involves the combined effect of dry bulb, wet bulb, and air movement as judged by the subjects in the research studies. There were a number of different combinations of dry bulb and relative humidity which would give the same feeling of comfort to a high percentage of the subjects for a given air velocity.

A typical comfort chart shown in Figure 4.1 could then be constructed by drawing lines through the points at which the majority of people equally clothed and equally active reported the same feeling of comfort. These lines are called the *effective temperature* (ET) lines.

The range of the summer effective temperatures from around 19 to 24°C, while the range of the effective temperatures are from 17°C to 22°C.

Over the years a number of similar charts have been developed by ASHRAE and other researchers including Fanger who developed General Comfort Charts based on clothing, activities, air temperatures, etc.

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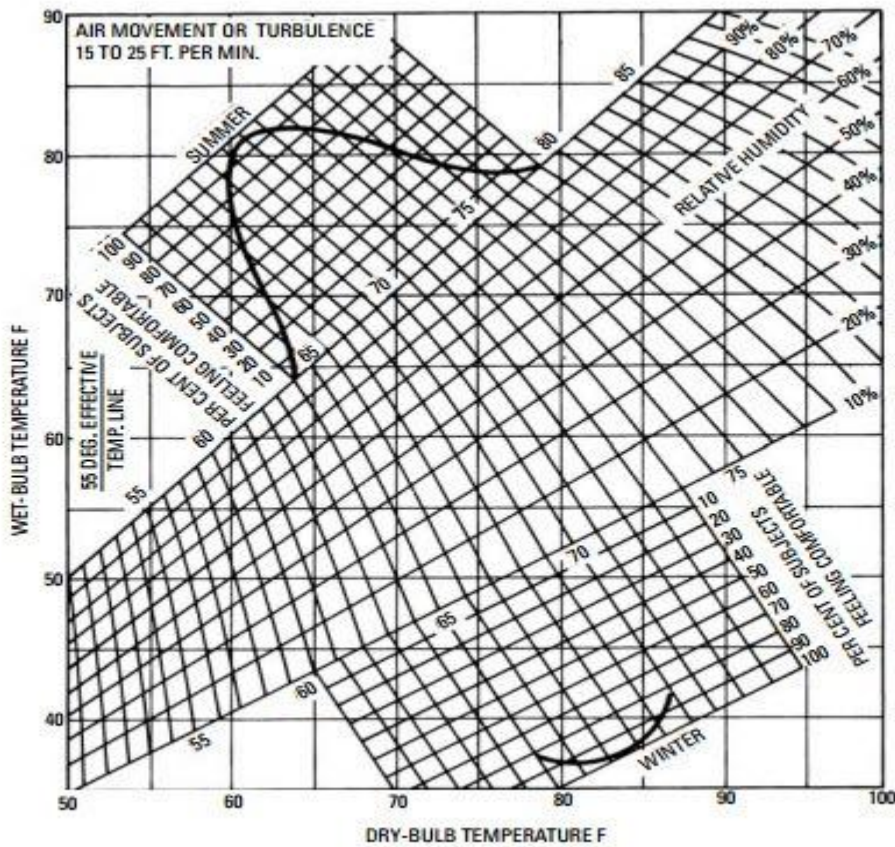
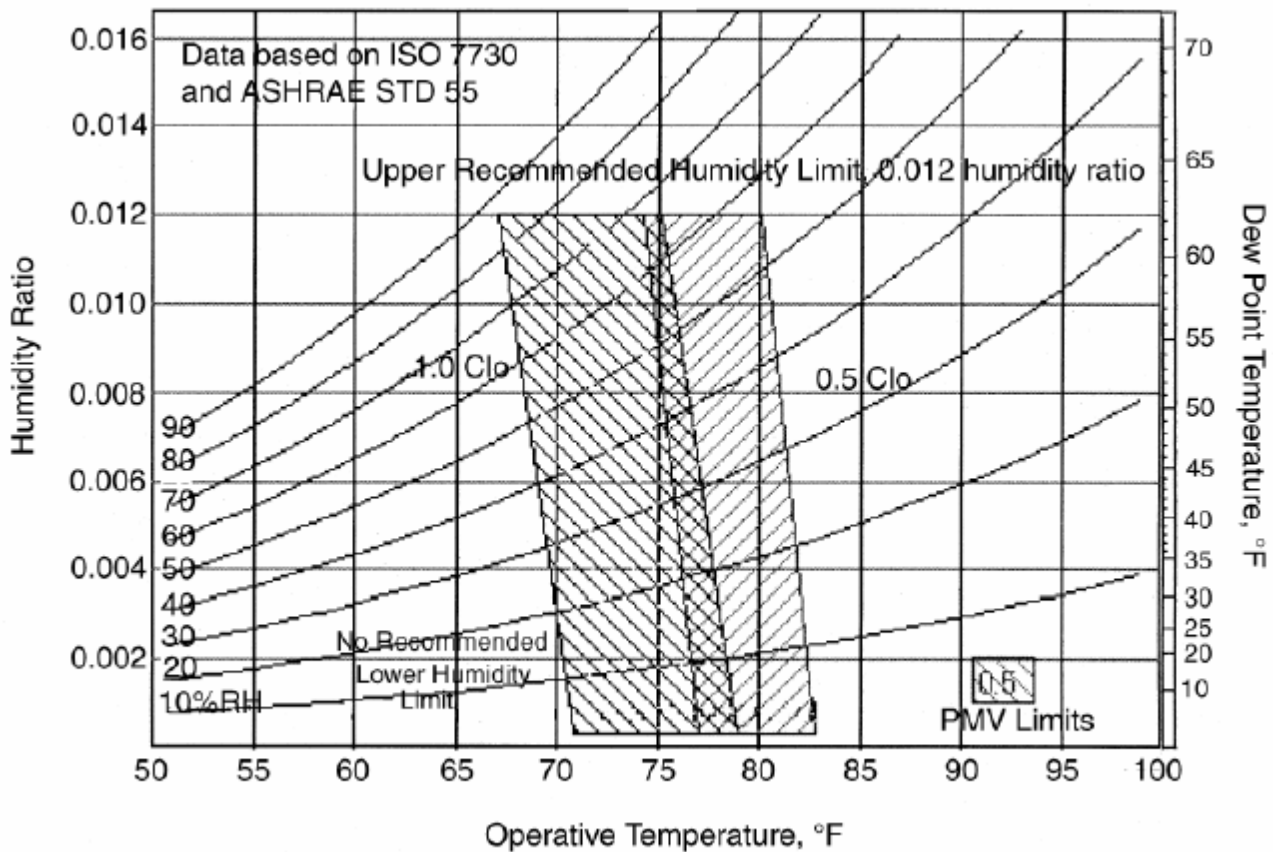


Figure 8-8 Comfort chart for still air. (Courtesy ASHRAE 1960 Guide)

(a) Comfort chart for still air

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(b) ASHRAE comfort zones for summer and winter

Figure 4.1 (a) Comfort chart for still air (b) ASHRAE comfort zones for summer and winter.

4.5 Prediction of Thermal Comfort

Human thermal is influenced by physiological factors, it is difficult to specify a single quantity for evaluating human comfort. The usual comfort parameters are ambient air temperature, humidity, air motion, body activity level, and clothing. However, it has been observed that if the surrounding surfaces are below the air dry bulb temperature, comfort would occur at a higher effective temperature than that indicated by Fig. 4.1. This implies that radiant cooling affects comfort parameters/sensation appreciably. Studies have also indicated that women of all ages prefer an effective temperature about one degree higher than that preferred by men, while both men and women over 40 years age prefer an effective temperature about one degree higher than that desired by younger people. People of all climatic regions have identical preferred temperatures. The activity level of the occupants and the duration of occupancy also affect human thermal comfort sensation.

Thermal comfort and thermal sensation can be predicted by (a) a comfort chart and (b) numerically by the predicted mean vote (**PMV**) and the predicted percentage of dissatisfied

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(PPD). The predicted mean vote predicts the mean response of a large group of people. This comfort sensation scale as developed by Rohles and Nevins is shown in Table 4.3.

Table 4.3 ASHRAE thermal sensation scale

+3	Hot
+2	Warm
+1	Slightly warm
0	Neutral
-1	Slightly cool
-2	Cool
-3	Cold

The test results have been correlated with the dry bulb temperature, humidity level, sex, and duration of exposure. The basic equation used to compute the **PMV** is

$$PMV = a^* t + b^* p_v + c^* \dots\dots\dots (4.4)$$

Where, **t** is the dry bulb temperature (°C) and **p_v** is the corresponding saturation pressure (kPa). **a^{*}**, **b^{*}** and **c^{*}** are the coefficients used for calculating PMV. The values of **a^{*}**, **b^{*}** and **c^{*}**, can be obtained from Table 4.4.

Table 4.4 Coefficient **a^{*}**, **b^{*}** and **c^{*}** used to calculate the predicted mean vote (PMV)

Exposure period (hr)	Sex	a [*]	b [*]	c [*]
1.0	Male	0.220	0.233	-5.673
1.0	Female	0.272	0.248	-7.245
1.0	Combined	0.245	0.248	-6.475
3.0	Male	0.212	0.293	-5.949
3.0	Female	0.275	0.255	-8.622
3.0	Combined	0.243	0.278	-6.802

For young adult subjects with sedentary activity and wearing clothing with a thermal resistance of approximately 0.5 clo, air velocity (0.2 m/s).

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After calculating the PMV, the PPD is estimated for the same condition (Fig. 4.2). The dissatisfied occupants are defined as those who do not vote either +1, 0 or -1 on the PMV scale. The PMV- PPD model is widely used and accepted for design and field assessment of comfort conditions.

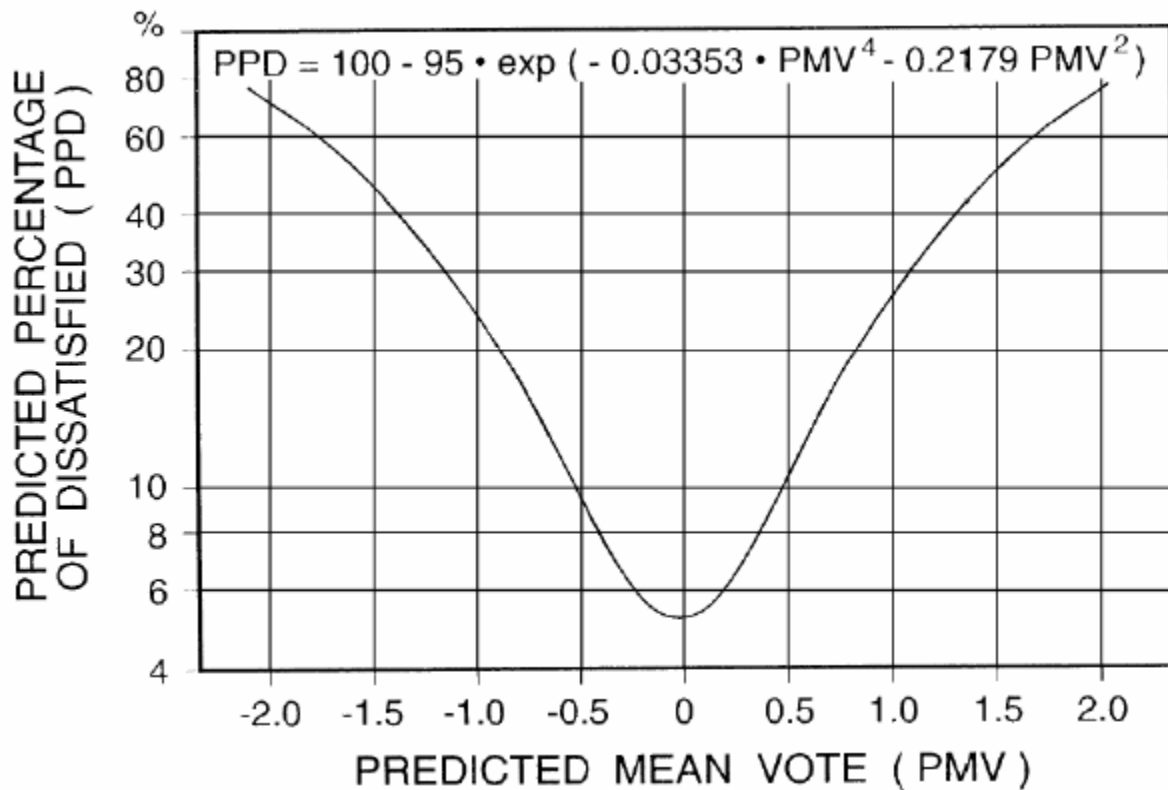


Figure 4.2 PPD as a function of PMV

Example 4.1:

A number of male and female subjects took part in a climate chamber test. Determine the difference between the PMV of male and that of female occupants with the dry bulb temperature being 24°C and the dew point temperature being 20°C, one hour after entry into the space.

Solution:

$$PMV = a^* t + b^* p_v + c^*$$

PMV for men:

$$PMV = 0.22(24) + 0.233(2.339) - 5.673 = 0.051$$

Similarly, PMV for women:

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$$PMV = 0.272(24) + 0.248(2.339) - 7.245 = - 0.136$$

From the above, both males and females are predicted to be thermally neutral.

The P_v values can be obtained from the table of thermodynamic properties of water at saturation.

4.6 Indoor Design Conditions

The most commonly recommended design conditions for comfort are:

$$ET^* = 24^\circ\text{C}$$

Dry bulb temperature = mean radiant temperature

Relative humidity = 50% (30 – 70)%

Air velocity less than 0.2 m/s

The indoor conditions to be maintained within a building are the dry bulb temperature and relative humidity of the air at the breathing line, 1 to 1.5 m above the floor, in an area that would indicate average condition at that level and which would not be affected by abnormal or unusual heat gains or losses from the interior or exterior.

Table 4.5 shows the guideline room air temperatures for different applications.

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Table 4.5

GUIDELINE ROOM AIR TEMPERATURES

Type of Space	°F		°C	
	Summer	Winter	Summer	Winter
Residences, apartments, hotel and motel guest rooms, convalescent homes, offices, conference rooms, classrooms, courtrooms, and hospital patient rooms	74–78	68–72	23–26	20–22
Theaters, auditoriums, churches, chapels, synagogues, assembly halls, lobbies, and lounges	76–80	70–72	24–27	21–22
Restaurants, cafeterias, and bars	72–78	68–70	22–26	20–21
School dining and lunch rooms	75–78	65–70	24–26	18–21
Ballrooms and dance halls	70–72	65–70	21–22	18–21
Retail shops and supermarkets	74–80	65–68	23–27	18–20
Medical operating rooms ^a	68–76	68–76	20–24	20–24
Medical delivery rooms ^a	70–76	70–76	21–24	21–24
Medical recovery rooms and nursery units	75	75	24	24
Medical intensive care rooms ^a	72–78	72–78	22–26	22–26
Special medical care nursery units ^a	75–80	75–80	24–27	24–27
Kitchens and laundries	76–80	65–68	24–27	18–20
Toilet rooms, service rooms, and corridors	80	68	27	20
Bathrooms and shower areas	75–80	70–75	24–27	21–24
Steam baths	110	110	43	43
Warm air baths	120	120	49	49
Gymnasiums and exercise rooms	68–72	55–65	20–22	13–18
Swimming pools	75 or above	75	24	24
Locker rooms	75–80	65–68	24–27	18–20
Children's play rooms	75–78	60–65	24–26	16–18
Factories and industrial shops	80–85	65–68	27–29	18–20
Machinery spaces, foundries, boiler shops, and garages	—	50–60	—	10–16
Industrial paint shops	—	75–80	—	24–27

^a Variable temperature range required with individual room control.

4.7 Quality and Quantity of Air

The air in an occupied space should, at all times, be free from toxic, unhealthy or disagreeable fumes such as carbon dioxide. It should also be free from dust and odors. In order to obtain these conditions, enough clean outside air must always be supplied to an occupied space to counteract or adequately dilute the source of contamination. The concentration of odor in a room depends upon many factors such as dietary and hygienic habits of occupants, type and amount of outdoor air supplied, room volume and type of odor sources. In general, when there is no smoking in a room, 1 m³/min per person of outside air will take care of all the conditions. But when smoking takes place in a room, 1.5 m³/min per person of outside air is necessary. For general application, a minimum of 0.3 m³/min of outside air per person, mixed with 0.6 m³/min of recirculate air is good.

Table 4.6 lists the ventilation standards while Table 4.7 gives the recommended volume of ventilation air for various buildings.

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Table 4.7 Ventilation Standards

Application	Smoking	m ³ /min per person		m ³ per m ² of floor area
		recommended	minimum	minimum
Apartment	Some	0.57	0.43	----
Banking space	Occasional	0.29	0.21	----
Barber shops	Considerable	0.43	0.28	----
Broker's board rooms	Very heavy	1.42	0.85	----
Corridors	----	----	----	0.007
Departmental stores	None	0.21	0.14	0.001
Director's rooms	Extreme	1.42	0.84	----
Drug stores	Considerable	0.29	0.21	----
Factories	None	0.29	0.21	0.003
Garage	----	----	----	0.028
Hospital operation	None	----	----	0.056
Room	None	0.85	0.71	0.009
Hospital private room	None	0.57	0.43	----
Hospital wards	Heavy	0.85	0.71	0.009
Hotel rooms	----	----	----	0.112
Kitchen restaurant	----	----	----	0.056
Kitchen residence	Heavy	1.42	----	0.056
Meeting rooms	None	----	----	----
School room	Some	0.43	0.28	----
Office general	None	0.71	0.43	0.007
Office private	----	----	----	----

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Table 4.8 Recommended volume of ventilation air for various buildings

Type of building	No. of changes per hour		m ³ /h per occupant		m ³ /m ² floor area per hour	
	min	Mix	min	Mix	min	mix
Commercial						
Garages	6	12	----	----	----	----
Offices	1.5	12	----	----	----	----
Waiting rooms	4	6	----	----	----	----
Restaurant (dining)	4	20	----	----	----	----
Restaurant (kitchen)	4	60	----	----	80	80
Stores	6	12	----	----	----	----
Farm Building						
Cow	----	----	100	100	----	----
Horse	----	----	120	140	----	----
Sheeps	----	----	200	300	----	----
Hospitals						
Dining rooms	6	12	----	----	30	30
Kitchens	20	60	----	----	80	80
Operating room	----	----	100	100	----	----
Toilets	75	30	----	----	40	40
Wards	----	----	70	150	20	20
Hostels						
Cafes	7.5	7.5	----	----	----	----
Dining rooms	4	20	----	----	30	30
Guest rooms	3	5	----	----	----	----
Kitchens	4	60	----	----	80	80
Lobbies	3	4	----	----	----	----
Toilets	10	12	----	----	----	----
Residences						
Bathrooms	1	5	----	----	----	----
Halls	1	3	----	----	----	----
Kitchen	1	40	----	----	----	----
Sleeping room	----	1	----	----	----	----
Various public spaces						
Auditorium, church, dance	4	30	20	140	30	40
Hall	----	----	60	80	40	40
Classrooms	12	12	----	----	30	30
Gymnasium	6	20	----	----	----	----
Laboratories	2	10	----	----	40	40
Locker rooms	30	30	----	----	40	40
Projection booths	3	5	----	----	----	----
Reading rooms	3	12	----	----	----	----
Engine and boiler room	----	----	----	----	----	----

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4.8 Outside Design Conditions

It is observed that there is a kind of sinusoidal relationship between the air dry bulb temperature and the sun time. For example, in the month of June in a certain locality where the sunrise is at about 5 a.m. and sunset at about 7 p.m. the time minimum temperature falls at about 4 a.m. and that of maximum temperature at about 4 p.m. i.e., with a lapse of about 12 hours. As regards relative humidity, it is seen that it reaches a minimum value in the afternoon. Since the mean daily maximum dry bulb temperature occurs between 1 p.m. and 4 p.m., it is reasonable to assume that the minimum relative humidity would occur during the same period.

The air conditioning load is estimated to become a base for selecting the conditioning equipment. It must take into account the entering into the space from outdoors on a design day, as well as the heat being generated within the space. The outside design data is available in a tabular form. Many guide and data books like ASHRAE give outside design conditions for different places this can be observed in Table 4.9.

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Heating Load Calculation

Heating loads are the thermal energy that must be supplied to the interior of a building in order to maintain the desired comfort conditions. There are two kinds of heat losses; (1) the heat transmitted through the walls, ceiling, floor, glass, or other surfaces, and (2) the heat required to warm outdoor air entering the space.

The actual heat loss problem is transient because the outdoor temperature, wind velocity, and sunlight are constantly changing, but during the coldest months, however sustained periods of very cold, cloudy, and stormy weather with relatively small variation in outdoor temperature may occur. In this situation heat loss from the space will be relatively constant and maximum. Therefore for design purposes the heat loss is usually estimated for steady-state heat transfer for some reasonable design temperature.

Heat loss (gain) calculations by the (BTU) or (kW) method is an accurate process of determining the heat transmission through building materials. The established (U) factors of combinations of building materials gives a simple corrected method of determining the loss of heat from a building.

With the cost of material continually rising, accurate heat loss is a must. Systems can be designed to do the job efficiently without the long used *safety factor*. The safety factor is only a "cost more" factor.

Conductivity for Plan Wall and Steady State

Fourier equation;

$$q = - KA \frac{dT}{dx} \rightarrow q = - \frac{dT}{(dx/KA)} \dots\dots\dots (5.1)$$

Where:

q: heat flow (kW)

dT: thermal potential difference (°C)

dx: wall thickness (m)

K: thermal conductivity (W/m.°C)

A: perpendicular area (m²)

We can say that; $R_{th} = dx/KA$

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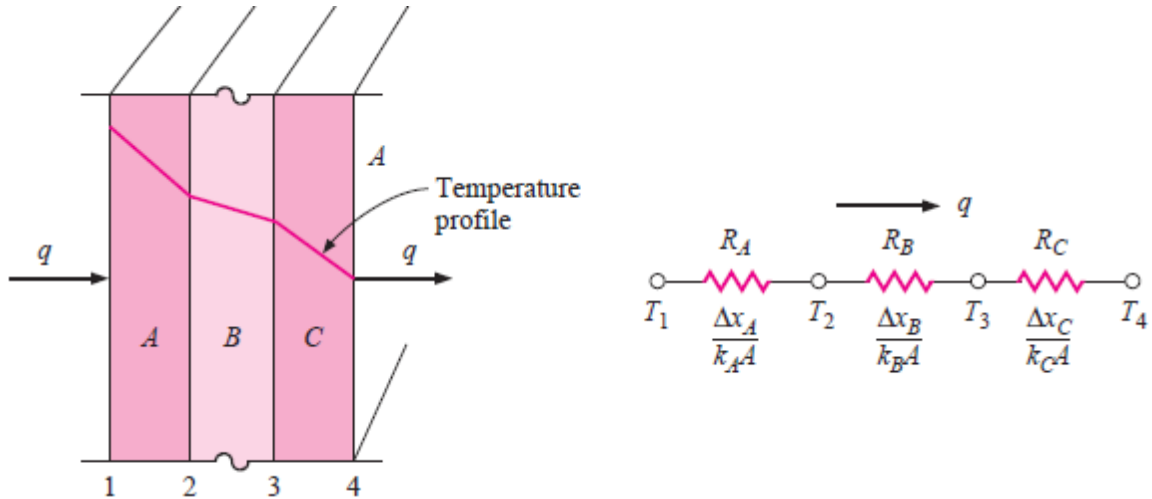


Figure 5.1 One-dimensional heat transfer through a composite wall and electrical analogy.

The temperature gradients in the three materials are shown in the Fig.5.1, and the heat flow may be written

$$q = - K_A A (T_2 - T_1) / \Delta x_A = - K_B A (T_3 - T_2) / \Delta x_B = - K_C A (T_4 - T_3) / \Delta x_C \dots\dots\dots (5.2)$$

$$q = (T_1 - T_4) / (\Delta x_A / K_A A + \Delta x_B / K_B A + \Delta x_C / K_C A)$$

$$= \Delta T_{\text{overall}} / \Sigma R_{\text{th}} \dots\dots\dots (5.3)$$

Convection heat transfer

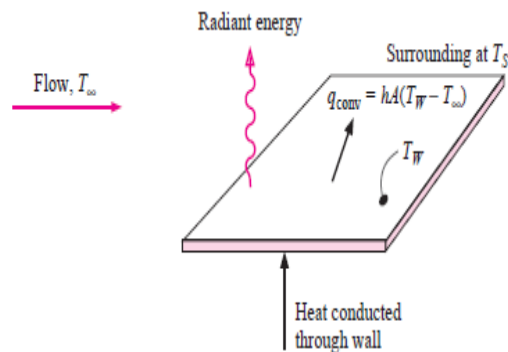


Figure 5.2 Convection heat transfer.

$$q = hA(T_w - T_\infty) \dots\dots\dots (5.4)$$

where; A is the surface area

$q = \Delta T / (1/h A)$, where $(1/h A)$ is the convection thermal resistance.

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Overall heat transfer coefficient

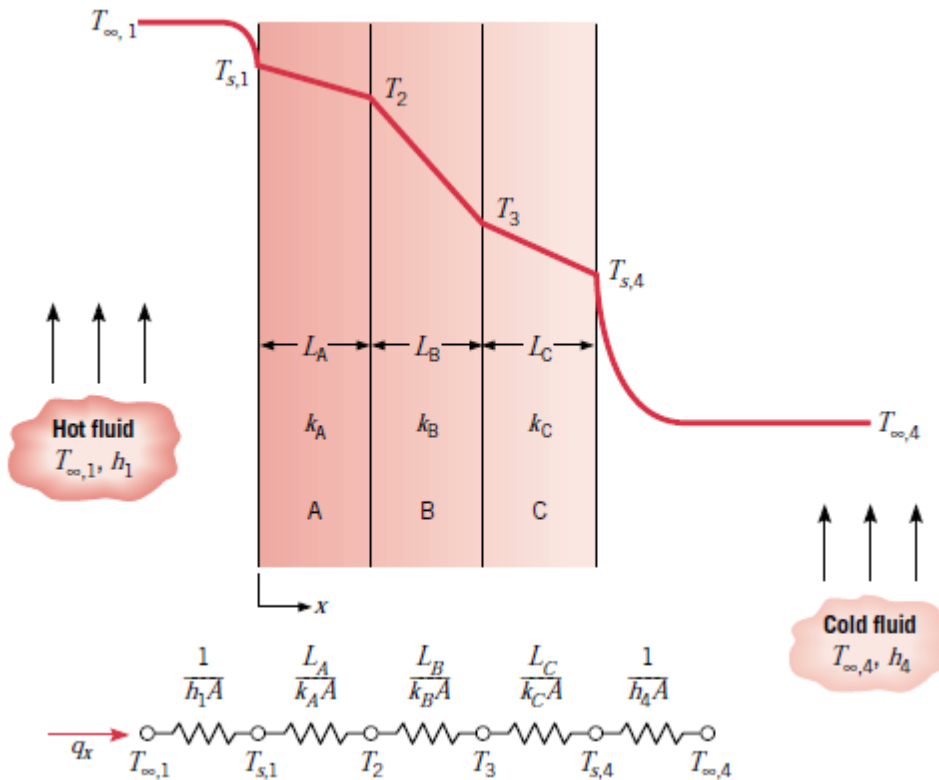


Figure 5.3 Equivalent thermal circuit for series composite wall.

$$q = \Delta T_{\text{overall}} / \Sigma R_{\text{th}}$$

$$\Delta T_{\text{overall}} = T_i - T_o$$

$$\Sigma R_{\text{th}} = R_i + R_A + R_B + R_C + R_o$$

Equation of heat transfer can be written as;

$$q = U A \Delta T_{\text{overall}} \dots \dots \dots (5.5)$$

Where

U is the overall heat transfer coefficient = $1 / \Sigma R_{\text{th}}$

To determine how much heat is lost through the walls, we have "U" factors for different types of construction. A "U" factor tells us how many (kW's) are transmitted to the colder (hotter) outside for one square meter of area with a difference of one degree between the outside surface and inside surface of the walls, windows, roofs or ceilings.

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Definitions

1. "U" – Coefficient of heat transmission (overall): The amount of heat transmitted from air to air per square meter of the wall, roof, or ceiling for a difference in temperature of one degree centigrade between the air on the inside and outside (winter) of the wall, floor, roof or ceiling.
2. "K" – Conductivity: The amount of heat (W) transmitted through one square meter of a homogeneous material one meter thick for a difference in temperature of one degree centigrade between two surfaces of the material.
3. "C" – Conductance: The amount of heat (W) transmitted from surface to surface through one square meter of material or construction for a difference in temperature of one degree centigrade between the two surfaces. This is not per meter or centimeter of thickness, but for thickness shown.
4. "f_o" – Outside film coefficient: The outside combined surface loss due to radiation and convection, with a 5 or 10 km/h (or more) wind velocity. The amount of heat (W) for one square meter of surface for a temperature difference of one degree centigrade (W/m².°C).
5. "f_i" – Inside film coefficient: The combined inside surface loss due to radiation and convection, with still air (W/m².°C).
6. "a" – Thermal conductance of an air space: The amount of heat (W) transmitted through one square meter of surface for a difference in temperature of one degree centigrade (W/m².°C).
7. "R" – This is the reciprocal of conductivity and conductance, or the overall heat transfer coefficient "U".

The total resistance "R" to heat flow through a wall is equal numerically to the sum of the resistance in series.

$$U_{\text{Tot}} = 1/R_{\text{Tot}}$$

$$R = 1/C, \quad R = 1/a, \quad R = 1/f, \quad R = X/K$$

$$U_{\text{Tot}} = 1/\Sigma R = 1/(1/f_i + 1/a + X_1/K_1 + X_2/K_2 + \dots + 1/f_o)$$

Example:

A wall consist of 10 cm face brick with 15 cm concrete and 1.3 cm cement plaster. The outdoor temperature is -7°C, and the wind velocity is 24 km/h, while the indoor space temperature is 25°C. Calculate:

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(a) Thermal resistance of the wall (b) Overall heat transfer coefficient

(c) Rate of heat transfer through the wall (d) Rate of heat transfer when neglect the thermal resistance of the air layers.

Solution:

Use the table of thermal conductivity of building materials to find the following:

For face brick $K_b = 1.30 \text{ W/m} \cdot ^\circ\text{C}$

For concrete $K_c = 1.73 \text{ W/m} \cdot ^\circ\text{C}$

For cement plaster $K_p = 0.721 \text{ W/m} \cdot ^\circ\text{C}$

$$f_i = 8.29 \text{ W/m}^2 \cdot ^\circ\text{C}, \quad f_o = 34.1 \text{ W/m}^2 \cdot ^\circ\text{C}$$

(a) The total thermal resistance (ΣR_{th})

$$\begin{aligned}\Sigma R_{th} &= 1/f_o + X_b/K_b + X_c/K_c + X_p/K_p + 1/f_i \\ &= 1/34.1 + 0.1/1.30 + 0.15/1.73 + 0.013/0.721 + 1/8.29 \\ &= 0.332 \text{ m}^2 \cdot ^\circ\text{C/W}\end{aligned}$$

(b) The overall heat transfer coefficient (U)

$$U = 1/\Sigma R_{th} = 1/0.332 = 3.012 \text{ W/m}^2 \cdot ^\circ\text{C}$$

(c) The rate of heat transfer through the wall (q/A)

$$q/A = U (T_i - T_o) = (3.012) (25 - (-7)) = 96.384 \text{ W/m}^2$$

$$\begin{aligned}\Sigma R_{th} &= X_b/K_b + X_c/K_c + X_p/K_p \\ &= 0.1/1.30 + 0.15/1.73 + 0.013/0.721 = 0.182 \text{ m}^2 \cdot ^\circ\text{C/W}\end{aligned}$$

$$q/A = (1/0.182) (25 + 7) = 175.82 \text{ W/m}^2$$

Example:

Door made of wood sheet and contains a glass in the middle, which represent 80% of the door area. The thickness of the door is 40 mm and the room temperature is 22°C , while the outdoor temperature is 5°C . Determine the rate of heat loss from the room, if the dimensions of the door are 2 m x 1 m.

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Solution:

$$\text{Door area} = 2 \times 1 = 2 \text{ m}^2$$

$$\text{Wood area} = 0.2 \times 2 = 0.4 \text{ m}^2$$

$$\text{Glass area} = 0.8 \times 2 = 1.6 \text{ m}^2$$

$$K_w = 0.159 \text{ W/m} \cdot ^\circ\text{C}, \quad C_g^* = 6.25 \text{ W/m}^2 \cdot ^\circ\text{C}$$

Rate of heat transfer through the wood of the door:

$$\begin{aligned} \Sigma R_{th} &= 1/f_o + X_w/K_w + 1/f_i \\ &= 0.121 + (0.04/0.159) + 0.029 = 0.402 \text{ m}^2\text{C/W} \end{aligned}$$

$$q_w = (1/0.402) \times 0.4 \times (22 - 5) = 16.9 \text{ W}$$

Rate of heat transfer through the glass:

$$q_g = U \cdot A \cdot (t_i - t_o) = 6.25 \times 1.6 \times (22 - 5) = 170 \text{ W}$$

Then the total heat loss through the door:

$$q_d = q_w + q_g = 16.9 + 170 = 186.9 \text{ W}.$$

The Temperature of the Surface of the Wall

It is important to know that the inside and outside air temperature is not equal to wall surface temperature. So in case when the temperature of inside wall surface reaches or below the dew point of indoor space temperature, it will causes to condensate the water vapor and may be defect this wall (especially in the winter). It is important to make a chick about this point.

Example:

An external wall consist of three layers, layer A ($X= 5 \text{ cm}$, $K= 0.4 \text{ W/m.K}$), layer B ($X= 24 \text{ cm}$, $K= 0.6 \text{ W/m.K}$), layer C ($X= 5 \text{ cm}$, $K= 0.8 \text{ W/m.K}$). If the indoor space condition is 20°C DBT and 14°C WBT, and the outdoor temperature is -15°C . Chick is the vapor condensate on the inner surface of the wall or not?

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Solution:

We must find the temperature of the inner surface (T_w) and compare it with the dew point of the air space, which is 10.5°C

$$U = 1/\Sigma R_{th} = 1/(1/f_i + X_A/K_A + X_B/K_B + X_C/K_C + 1/f_o)$$
$$= 1/(1/8.29 + 0.05/0.4 + 0.24/0.6 + 0.05/0.8 + 1/34.1) = 1.36 \text{ W/m}^2.\text{C}$$

$$q = U.A.(T_i - T_o) = 1.36 \times 1 \times (20 + 15) = 47.6 \text{ W/m}^2$$

Also

$$q = f_i (T_i - T_s) \quad \text{where; } T_s \text{ is the inner surface temperature}$$

$$47.6 = 8.29 \times (20 - T_s) \rightarrow T_s = 14.26^\circ\text{C}$$

The inner surface temperature is greater than the dew point, so the vapor will not condenses.

Example:

A wall was constructed of hollow concrete blocks ($C = 5.11 \text{ W/m}^2.\text{C}$). The indoor condition is 20°C DBT and 14°C WBT, when the outdoor air temperature is -15°C and the wind velocity is 24 km/h in winter. Determine the inner surface temperature, and show if the vapor condenses on it or no.

Solution:

The dew point for the indoor air from the psychrometric chart is = 9.3°C

$$\Sigma R_{th} = 1/f_i + 1/C + 1/f_o = (1/8.29) + (1/5.11) + (1/34.1) = 0.346 \text{ m}^2\text{C/W}$$

$$R_f/R_t = \Delta T_f/\Delta T_t$$

$$0.121/0.346 = (20 - T_s)/(20 + 15)$$

$$T_s = 7.8^\circ\text{C}$$

The inner surface temperature is lower than the dew point, so the vapor will condenses.

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Infiltration

All structures have some air leakage or infiltration. This means a heat loss because the cold dry outdoor air must be heated to the inside design temperature and moisture must be added to increase the humidity to the design value. The heat required to increase the temperature is given by

$$q_s = m \cdot C_p (T_i - T_o) \dots\dots\dots (5.6)$$

Where

q_s : sensible heat loss

m : mass flow rate of the infiltration air

C_p : specific heat of moist air

Infiltration is usually estimated on the basis of volume flow rate at outdoor conditions.

Then equation 5.6 become:

$$q_s = (V/v) C_p (T_i - T_o) = 1.22 V (T_i - T_o) \dots\dots\dots (5.7)$$

Where

V : infiltration volume flow rate (m^3/s)

v : specific volume (m^3/kg)

The latent heat required to humidify the air is given by

$$q_L = m (W_i - W_o) h_{fg} \dots\dots\dots (5.8)$$

$$= (V/v) (W_i - W_o) h_{fg} = 3010 V (W_i - W_o) \dots\dots\dots (5.9)$$

Where

$(W_i - W_o)$: the difference in design moisture ratio ($kg/kg_{d.a}$)

h_{fg} : the latent heat of vaporization at indoor condition (J/kg) or (kJ/kg)

More than one method is used in estimating air infiltration in building structures.

1. Air Change Method

Experience and judgment are required to obtain satisfactory results with this method. Experienced engineers will often simple make an assumption of the number of air changes per hour (ACH) that a building will experience based on their appraisal of the building type,

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construction, and use. The range will usually be from 0.5 ACH (very low) to 2.0 ACH (very high). This approach is usually satisfactory for design load calculation but not recommended for the beginner.

In practice, the following values of air changes per hour can be used with reasonable precision for rooms with the extent of windows and external doors given.

No windows or exterior doors	0.5
Exterior doors or windows on one side	1
Exterior doors or windows on two sides	1.5
Exterior doors or windows on three sides	2
Entrance halls	2

2. Crack Method

The flow (leak) through an opening is proportional to the area of the cracks, the type of the cracks, and the pressure difference across the crack.

$$V = A.C.\Delta P^n \dots\dots\dots (5.10)$$

Where

A: effective leak area of the cracks

C: flow coefficient, which depends on the type of crack and the nature of the flow in the crack.

ΔP : outside – inside pressure difference ($P_o - P_i$)

N: Exponent that depends on the nature of the flow in the crack $0.4 < n < 1.0$.

The following table gives the leakage rates through cracks in doors on the windward side for different wind velocities and different door constructions.

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Type of door	m ³ per linear meter of crack					
	Wind velocity, km/h					
	8	16	24	32	40	48
Glass door, good installation						
3.2 mm crack	0.3	0.6	0.9	1.21	1.49	1.77
Average installation						
4.76 mm crack	0.45	0.93	1.3	1.86	2.23	2.7
Poor installation						
6.4 mm crack	0.6	1.21	1.77	2.42	2.42	3.53
Ordinary wood or metal door						
well fitted W-stripped	0.04	0.06	0.08	0.12	0.16	0.2
well fitted now W-stripped	0.08	0.11	0.17	0.24	0.31	0.39
Poorly fitted						
Now W-stripped	0.08	0.21	0.34	0.48	0.61	0.78
Factory door 3.2 mm crack	0.3	0.6	0.9	1.21	1.49	1.77

Example:

Find the sensible heat loss due to infiltration outdoor air at -6°C into a heated house through an aluminum window (2 m x 1.2 m), which has double swinging parts the width of each is 0.5 m, and a fixed glass in the middle. The temperature of the air inside the house is 22°C.

Solution:

$$\text{Crack length} = 2[(1.2 + 0.5) \times 2] = 6.8 \text{ m}$$

ASHRAE Handbook, 1981 gives value of 0.77 (l/s. m) infiltration air through aluminum windows when the air velocity is 40 km/h.

$$V = (0.77/1000) \times 6.8 = 0.00524 \text{ m}^3/\text{s}$$

And from equation 5.7

$$q_s = 1.22 V (T_i - T_o) = 1.22 (0.00524)(22 + 6) = 0.179 \text{ kW.}$$

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Ventilation

The introduction of outdoor air for ventilation of conditional spaces is necessary to dilute the odors given off by people, smoking and other internal air contaminates.

The amount of ventilation required varies primarily with the total number of people, the ceiling height and the number of people smoking. People give off body odors which required a minimum of 5 cfm (2.36 l/s) per person. When people smoke, the additional odors given off by cigarettes or cigars requires a minimum of 15 to 25 cfm (7 – 12 l/s) per person. In special gathering rooms with heavy smoking, 30 to 50 cfm (15 – 24 l/s) per person is recommended.

Tables 4.7 and 4.8 are used to determine the minimum and recommended ventilation air quantity for several applications.

The sensible and latent heating loads from ventilation air can be estimated by the following equations

$$q_s = (V/v) C_p (T_i - T_o) = 1.22 V \cdot (T_i - T_o) \dots\dots\dots (5.10)$$

$$q_L = (V/v) (W_i - W_o) h_{fg} = 3010 V \cdot (W_i - W_o) \dots\dots\dots (5.11)$$

V: ventilation volume flow rate (m³/s)

Example:

An auditorium seats 1000 people. The space design conditions are 21°C and 40% RH, and outdoor design conditions 5°C DBT and 60% RH. What is the heating load due to ventilation.

Solution:

From table 4.8 the minimum ventilation air per person to be 5.5 l/s

Total ventilation air = 5.5 x 1000 = 5500 l/s = 5.5 m³/s

$$q_s = (V/v) C_p (T_i - T_o) = 1.22 V \cdot (T_i - T_o) = 1.22 (5.5)(21 - 5)$$

$W_i = 0.0061$ kg/kg dry air , $W_o = 0.0034$ kg/kg dry air [from the psychrometric chart at the inner and outer design conditions]

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$$q_L = (V/v) (W_i - W_o) h_{fg} = 3010 V \cdot (W_i - W_o) \\ = 3010 (5.5)(0.0061 - 0.0034) = 44.7 \text{ kW}$$

Then the total heating load due to ventilation

$$Q_v = 107.36 + 44.7 = 152.06 \text{ kW}$$

Heat Losses from Air Ducts

The losses of a duct system can be considerable when the ducts are not in the conditioned space. Proper insulation will reduce these losses but cannot completely eliminate them. The losses may be estimated using the following relation

$$q_D = U \cdot A_s \cdot \Delta T_m \dots\dots\dots (5.12)$$

where

U: overall heat transfer coefficient (W/m²°C)

A_s: outside duct surface area (m²)

ΔT_m: mean temperature difference between the air in the duct and the environment.

Example:

Estimate the heat losses from 0.5 m³/s of air at 50°C round duct 8 m in length. The duct has 25 mm (1 in) of fibrous glass insulation and the overall heat transfer coefficient is 1.1357 W/m².°C. The environment temperature is -10°C and the duct diameter is 400 mm.

Solution:

Equation 5.12 will be used to estimate the heat losses

$$\Delta T_m = T_s - T_o = 50 - (-10) = 60^\circ\text{C}$$

The surface area of the duct is

$$A_s = \pi \cdot D \cdot L = \pi (0.4 + 0.05)(8) = 11.3 \text{ m}^2$$

$$\text{Then, } q_D = 1.1357 (11.3)(60) = 770 \text{ W} = 0.77 \text{ kW}$$

the temperature of air leaving the duct may be computed from

$$q = m \cdot C_p (T_2 - T_1) = V \cdot \rho \cdot C_p (T_2 - T_1)$$

$$T_2 = T_1 + q / (V \cdot \rho \cdot C_p) = 50 + (-0.77 / 0.5 \times 1.2 \times 1.017) = 48.7^\circ\text{C}$$

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Minimum insulation of supply and return duct is presently specified by ASHRAE standards as follows:

All duct system shall be insulated to provide a thermal resistance, excluding film resistance, of

$$R = \Delta T / 47.3 \text{ (m}^2\text{C/W)}$$

Where, $\Delta T = T_{\text{duct}} - T_{\text{surrounding}}$

Heat losses from the supply ducts become part of the space heating load and should be summed with transmission and infiltration heat losses.

Heat losses from the return ducts are not a part of the space heat loss, but should added to the heating equipment load.

Air Required for Space Heating

The air quantity is computed from

$$q_s = m \cdot C_p (T_s - T_r) = 1.22 V \cdot (T_s - T_r) \dots\dots\dots (5.13)$$

$$V = q_s / [1.22(T_s - T_r)]$$

Where

V : volume flow rate of supplied air, m^3/s

v : specific volume of supplied air, m^3/kg

T_s : temperature of supplied air, $^{\circ}\text{C}$

T_r : temperature of room (conditioned space), $^{\circ}\text{C}$

The temperature difference ($T_s - T_r$) is normally less than 10°C . It can be considered as a temperature difference about 6°C is suitable for most of the comfort air conditioning applications in Arab region.

After the total air flow rate required for the complete structure has been determine. The next step is to allocate the correct portion of the air to each room or space. Air quantity for each room should be apportioned according to the heating load for the space; therefore

$$V_{m} = V (q_{m}/q_t) \dots\dots\dots (5.14)$$

Where; q_m : total heat loss of room(n), W

V_m : volume flow rate of air supply to room(n), m^3/s

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Exercises:

1. The dimensions of the outer wall of a room are 6 m x 3 m, and it contains a glass window of dimensions 2 m x 1.5 m. The wall is made up of 24 cm brick with 10 cm face brick and 2 cm of a gypsum layer from the inside. The indoor air temperature is 20°C, while the outdoor air temperature and wind velocity are 5°C and 24 km/h respectively. Calculate:

(a) The total heat loss from the wall and the window together.

(b) The temperature of the inner surface of the wall.

2. An external wall is made of 100 mm common brick, with 40 mm gypsum plaster. What is the thickness of rock wool insulation ($K = 0.04 \text{ W/m}\cdot^\circ\text{C}$) to reduce the rate of heat loss by 80%.

3. The room conditions in winter are 22°C DBT and 40% RH. Infiltration air at 1°C DBT and 50% RH enters the room at a rate of 0.006 m³/s.

(a) Calculate the rate of water that must be evaporating to maintain the room conditions.

(b) Calculate the amount of heat required to vaporize the water at the above rate.

4. A building wall consists of 25 cm concrete ($K = 1.75 \text{ W/m}\cdot\text{K}$) and 1.9 cm plaster ($K = 87 \text{ W/m}\cdot\text{K}$) on the inside surface. The outside and inside surface heat-transfer coefficients can be taken as 34 and 9.4 W/m²K respectively. The outside temperature is -18°C, while the room is held at 23.5°C DBT and 16.8°C WBT.

(a) What is the temperature on the inside wall surface.

(b) Will the moisture condense on the wall.

(c) How many layers of 1.25 cm thick fiber-board insulation ($K = 0.048 \text{ W/m}\cdot\text{K}$) should be applied on the inside wall surface to prevent moisture to condense on it.

5. A residential house contains a reception room and a family room and two-bedroom. The heating load of the listed rooms are; 5.5 kW, 3.5 kW and 4.2 kW for each bedroom respectively. The inside required temperature is 21°C. Find the rates of hot air which required for every room.